

Quick Lab

Purpose To use a graph to identify period and group trends

Materials

- graph paper
- pencil



Periodic Trends in Ionic Radii

Procedure

Use the data presented in Figure 6.23 to plot ionic radius versus atomic number.

Analyze and Conclude

- Compare** How does the size change when an atom forms a cation and when an atom forms an anion?
- Describe** How do the ionic radii vary within a group of metals? How do they vary within a group of nonmetals?
- Describe** What is the shape of a portion of the graph that corresponds to one period?
- Compare and Contrast** Is the trend across a period similar or different for Periods 2, 3, 4, and 5?
- Explain** Propose explanations for the trends you have described for ionic radii within groups and across periods.

Figure 6.23 Atomic and Ionic Radii
Atomic and ionic radii are an indication of the relative size of atoms and ions. The data listed are reported in picometers (pm).



Periodic Trends In Ionic Radii Lab Answers

CL Gary



Periodic Trends In Ionic Radii Lab Answers:

Whispering the Techniques of Language: An Mental Journey through **Periodic Trends In Ionic Radii Lab Answers**

In a digitally-driven world where monitors reign great and quick transmission drowns out the subtleties of language, the profound strategies and mental subtleties hidden within phrases often get unheard. Yet, nestled within the pages of **Periodic Trends In Ionic Radii Lab Answers** a fascinating literary prize blinking with natural emotions, lies a fantastic journey waiting to be undertaken. Penned by an experienced wordsmith, that enchanting opus encourages readers on an introspective journey, delicately unraveling the veiled truths and profound influence resonating within the cloth of each and every word. Within the psychological depths of this emotional evaluation, we shall embark upon a heartfelt exploration of the book is key subjects, dissect its captivating publishing design, and yield to the effective resonance it evokes deep within the recesses of readers hearts.

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