CHAPTER 6 STUDY GUIDE

The Periodic Table and Periodic Law

Section 6.1 Development of the Modern Periodic Table

In your textbook, reads about the history of the periodic table's development.

Use each of the terms below just once to complete the passage.

octaves	atomic mass	atomia number	-ning-
colometer	properties .	Henry Messley	
protons	percodic law	Benissi Mandaloov	accepted

The table below was developed by John Newlands and is based on a relationship called the law of (1) Properties According to this law, the properties of the elements repeated every (2) eight elements. Thus, for example, element two and element (3) nine have similar properties. The law of octaves did not work for all the known elements and was not generally (4) OCCEPTED

1	1 2	3	4	5	6	7
н	Li	G	Bo	C	N	0
8	9	10	TI	12	13	1.5
· E	Na	Mg	- Al	Si	P	S

The first periodic table is mostly credited to (5) Mendelety In his table, the elements were arranged according to increasing (6) Office Mass. One important result of this table was that the existence and properties of undiscovered

(7) Elements could be predicted.

The element in the modern periodic table are arranged according to increasing

(8) Offeric number as a result of the work of (9) Moseloy . This arrangement is based on number of (10) profess in the nucleus of an atom of the element. The modern form of the periodic table results in the (11) periodic last. which states that when elements are arranged according to increasing atomic number, there is a periodic repetition of their chemical and physical (12) properties.

Periodic Table And Periodic Law Solution Manual

Arturo Cuomo

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