

Chapter 1: Limiting Reagents and Percent Yield (Pages 1-4)

Remember: Limiting reagent is the reactant that will completely run out first in a chemical reaction.
Calculate: Moles of product formed from a limiting reagent using stoichiometry. Percent yield = (actual yield / theoretical yield) x 100.

Chemical reactions involve the breaking and forming of chemical bonds. In a chemical reaction, the atoms of the reactants are rearranged to form the products.



The reaction of hydrogen and oxygen to form water is an example of a chemical reaction.

2H ₂ (g)	100.00 grams
O ₂ (g)	50.00 grams
H ₂ O(l)	100.00 grams
H ₂ (g)	100.00 grams

1. 100.00 g of hydrogen and 50.00 g of oxygen react to form water. What is the limiting reagent? What is the percent yield of water?



2. 100.00 g of hydrogen and 50.00 g of oxygen react to form water. What is the limiting reagent?

3. 100.00 g of hydrogen and 50.00 g of oxygen react to form water. What is the percent yield of water?



4. 100.00 g of hydrogen and 50.00 g of oxygen react to form water. What is the percent yield of water?



Percent Yield And Limiting Reactants Quiz

N Colangelo



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