

Percent Composition and Molecular Formula Worksheet Key

1. What's the empirical formula of a molecule containing 65.5% carbon, 5.5% hydrogen, and 29.0% oxygen? **C_3H_3O mass = 55 g/mole**
2. If the molar mass of the compound in problem 1 is 110 grams/mole, what's the molecular formula? **$C_6H_6O_2$**
3. What's the empirical formula of a molecule containing 18.7% lithium, 16.3% carbon, and 65.0% oxygen? **Li_2CO_3**
4. If the molar mass of the compound in problem 3 is 73.8 grams/mole, what's the molecular formula? **Li_2CO_3**

Write the molecular formulas of the following compounds:

5. A compound with an empirical formula of C_2OH_4 and a molar mass of 88 grams per mole. **$C_4O_2H_8$**
6. A compound with an empirical formula of C_4H_4O and a molar mass of 136 grams per mole. **$C_8H_8O_2$**
7. A compound with an empirical formula of $CFBrO$ and a molar mass of 254.7 grams per mole. **$C_2F_2Br_2O_2$**
8. A compound with an empirical formula of C_2H_8N and a molar mass of 46 grams per mole. **C_2H_8N**

Answer the following questions:

9. The percentage composition of acetic acid is found to be 39.9% C, 6.7% H, and 53.4% O. Determine the empirical formula of acetic acid. **CH_2O**
10. The molar mass for question #9 was determined by experiment to be 60.0 g/mol. What is the molecular formula? **$C_2H_4O_2$**
11. Aniline, a starting material for urethane plastic foams, consists of C, H, and N. Combustion of such compounds yields CO_2 , H_2O , and N_2 as products. If the combustion of 9.71 g of aniline yields 6.63 g H_2O and 1.46 g N_2 , what is its empirical formula? **C_6H_7N**
12. The molar mass of aniline is 93 g/mol. What is its molecular formula? **C_6H_7N**
13. Calculate the mass percent of carbon, nitrogen and oxygen in acetamide, C_2H_5NO . **%C 40.668 %H 8.533 %N 23.713 %O 27.086**
14. A 50.51 g sample of a compound made from phosphorus and chlorine is decomposed. Analysis of the products showed that 11.39 g of phosphorus atoms were produced. What is the empirical formula of the compound? **PCl_3**
15. When 2.5000 g of an oxide of mercury, (Hg_xO_y) is decomposed into the elements by heating, 2.405 g of mercury are produced. Calculate the empirical formula. **Hg_2O**
16. The compound benzamide has the following percent composition. What is the empirical formula?
C = 69.40 % H= 5.825 % O = 13.21 % N= 11.57 % **C_7H_7NO**
17. A component of protein called serine has an approximate molar mass of 100 g/mole. If the percent composition is as follows, what is the empirical and molecular formula of serine?
C = 34.95 % H= 6.844 % O = 46.56 % N= 13.59 %

$C_3H_7NO_3$ empirical formula

$C_3H_7NO_3$ molecular formula

Percent Composition And Molecular Formula Answers

Morris Hein, Susan Arena



Percent Composition And Molecular Formula Answers:

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Hren, 2014-11-24 Hundreds of practice problems to help you conquer chemistry Are you confounded by chemistry Subject by subject problem by problem Chemistry Workbook For Dummies lends a helping hand so you can make sense of this often intimidating subject Packed with hundreds of practice problems that cover the gamut of everything you ll encounter in your introductory chemistry course this hands on guide will have you working your way through basic chemistry in no time You can pick and choose the chapters and types of problems that challenge you the most or you can work from cover to cover With plenty of practice problems on everything from matter and molecules to moles and measurements Chemistry Workbook For Dummies has everything you need to score higher in chemistry Practice on hundreds of beginning to advanced chemistry problems Review key chemistry concepts Get complete answer explanations for all problems Focus on the exact topics of a typical introductory chemistry course If you re a chemistry student who gets lost halfway through a problem or worse yet doesn t know where to begin Chemistry Workbook For Dummies is packed with chemistry practice problems that will have

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this textbook and are available at no extra cost **Access to Chemistry** Avril Higton, Mike Clemmet, Elaine Golding, Alan V Jones, 2007-10-31 This text has been specifically designed to prepare people with previously limited chemical knowledge for entrance into science related courses such as Foundation and Access courses which involve chemistry in higher education Until now there have been no texts available for use on these courses and this book fills that gap Access to Chemistry effectively forms a self study course which is split into separate modules and units covering the full spread of concepts required for those needing a basic knowledge of chemistry The material is presented in a friendly and easy to use manner which allows the student to pace their acquisition of knowledge and gain increasing confidence in order to succeed in understanding essential relevant concepts Other useful features of this book include starter diagnostic tests worked examples and self study tests with answers at the end of each unit In addition to Access or Foundation course students and their tutors to whom this book will prove essential it will have an appeal also as a revision text for those needing a refresher after a break in the subject In addition it will be of interest to members of the general public who wish to better educate themselves on chemical matters as it provides a clear and useful insight into areas such as health home chemicals business market trends and gardening **E3 Chemistry Review Book - 2018 Home Edition (Answer Key Included)** Effiong Eyo, 2017-10-20 With Answer Key to All Questions Chemistry students and homeschoolers Go beyond just passing Enhance your understanding of chemistry and get higher marks on homework quizzes tests and the regents exam with E3 Chemistry Review Book 2018 With E3 Chemistry Review Book students will get clean clear engaging exciting and easy to understand high school chemistry concepts with emphasis on New York State Regents Chemistry the Physical Setting Easy to read format to help students easily remember key and must know chemistry materials Several example problems with solutions to study and follow Several practice multiple choice and short answer questions at the end of each lesson to test understanding of the materials 12 topics of Regents question sets and 3 most recent Regents exams to practice and prep for any Regents Exam This is the Home Edition of the book Also available in School Edition ISBN 978 197836229 The Home Edition contains an answer key section Teachers who want to recommend our Review Book to their students should recommend the Home Edition Students and and parents whose school is not using the Review Book as instructional material as well as homeschoolers should buy the Home Edition The School Edition does not have answer key in the book A separate answer key booklet is provided to teachers with a class order of the book Whether you are using the school or Home Edition our E3 Chemistry Review Book makes a great supplemental instructional and test prep resource that can be used from the beginning to the end of the school year PLEASE NOTE Although reading contents in both the school and home editions are identical there are slight differences in question numbers choices and pages between the two editions Students whose school is using the Review Book as instructional material SHOULD NOT buy the Home Edition Also available in paperback print

Understanding Essential Chemistry Max Diem, 2025-06-11 Enables students to understand apply and retain key concepts

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Foundations of College Chemistry Morris Hein, Susan Arena, Cary Willard, 2016-08-02 This text is an unbound three hole punched version Used by over 750 000 students Foundations of College Chemistry Binder Ready Version 15th Edition is praised for its accuracy clear no nonsense approach and direct writing style Foundations direct and straightforward explanations focus on problem solving making it the most dependable text on the market Its comprehensive scope proven track record outstanding in text examples and problem sets were all designed to provide instructors with a solid text while not overwhelming students in a difficult course Foundations fits into the prep intro chemistry courses which often include a wide mix of students from science majors not yet ready for general chemistry allied health students in their 1st semester of a GOB sequence science education students for elementary school teachers to the occasional liberal arts student fulfilling a science requirement Foundations was specifically designed to meet this wide array of needs

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