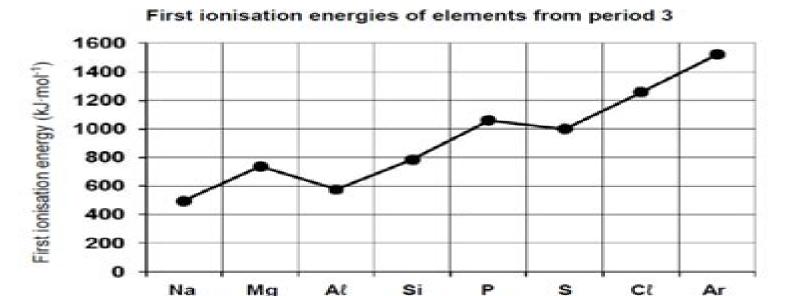
QUESTION 6

The graph below shows the first ionisation energies of elements from period 3 in the periodic table.



- 6.1 Define the term first ionisation energy.
- 6.2 Explain the difference in first ionisation energy between sodium and magnesium.
- 6.3 From the graph, write down the approximate first ionisation energy of sodium. (1)

Element

- 6.4 Explain why magnesium becomes ionised easily. (2)
- 6.5 Use the information in the graph above and calculate the energy needed to remove an electron from sodium atom. (3)
- 6.6 Draw an Aufbau diagram of a sodium ion. (2)
- 6.7 The Haber process uses the reaction between nitrogen and hydrogen to make ammonia. The formation of ammonia is exothermic, releasing 46 kilojoules per mole of ammonia formed.

$$3N_{2(g)} + H_{2(g)} \rightarrow 2NH_{3(g)}$$

- 6.7.1 What type of bond exists in the molecule of ammonia? (1)
- 6.7.2 Explain your answer to QUESTION 6.7.1.

(2)

(2)

(2)

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Xiang Xie

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Quantum Physics Stephen Gasiorowicz,2003-04-17 Since the publication of the first edition over 35 years ago Quantum Physics has been one of the standard quantum mechanics texts for undergraduate physics majors Its hallmarks are clear concise exposition and a balance of theory and applications In the 3rd Edition the author has made numerous changes based on feedback from teachers and students to enhance the book s strengths One of the author s important goals has been to offer teachers and students a textbook that is manageable in one semester However recognizing that some teachers like to go into more depth on certain topics he has developed a web site where more detailed presentations can be found

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