

7. Give a recipe for preparing a sodium carbonate/sodium bicarbonate buffer solution having a pH of 9.5. **For carbonic acid, H_2CO_3 , $\text{pK}_{\text{a}1} = 6.35$ and $\text{pK}_{\text{a}2} = 10.33$**

8. A saturated solution of silver chromate (Ag_2CrO_4 , a sparingly soluble solid, formula mass 331.0 g/mol), was found to have a chromate ion concentration of $3.2 \times 10^{-4} \text{ M}$ at 25°C . What is (a) the K_{sp} ; and (b) solubility of the salt in mg/mL? **How could you dissolve a sample of the salt?**

9. Calculate the solubility in mg/mL of PbCl_2 (formula mass 276.2 g/mol) in 0.25 M sodium chloride solution at 25°C . K_{sp} for lead chloride is 1.8×10^{-10} at 25°C .

10. Write an equation to show the formation of the complex ion $[\text{Ag}(\text{NH}_3)_2]^+$ from silver ion and aqueous ammonia then calculate the concentration of silver ions in solution if exactly 0.100 mol of silver nitrate is added to exactly **500 mL of 0.800 M ammonia** solution. K_f for $[\text{Ag}(\text{NH}_3)_2]^+$ is 1.7×10^7 .

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D Kirk



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