

# **10-Line Essay on Save the Environment**

- 1. The environment is important for all living things.**
- 2. We need clean air, water, and healthy land to live.**
- 3. Pollution harms the environment and makes life hard.**
- 4. Cutting down trees makes the air dirty and harms animals.**
- 5. Protecting nature helps make the world a better place.**
- 6. We can plant trees to keep the air fresh.**
- 7. Saving water is another way to protect the environment.**
- 8. Recycling helps reduce waste and keep nature clean.**
- 9. Kids can help by picking up trash and planting trees.**
- 10. Saving the environment is important for everyone.**

# Planting Trees Way Saving Environment Essay

**Chandranna NS**



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web chemistry 120 hess s law worksheet calculate  $\Delta h$  for the reaction  $\text{C}_2\text{H}_4(\text{g}) + \text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g})$  from the following data  $\text{C}_2\text{H}_4(\text{g}) + 3 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 2 \text{ H}_2\text{O}(\text{l})$   $\Delta h = 1411 \text{ kJ}$   $\text{C}_2\text{H}_6(\text{g}) + 7 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{l})$   $\Delta h = 1560 \text{ kJ}$   $\text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$   $\Delta h = 285 \text{ kJ}$

mole

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web thus for reaction 9 2 in which 1 mole of  $\text{H}_2\text{SO}_4$  and 1 mole of  $\text{NaOH}$  combine to form 2 moles of  $\text{H}_2\text{O}$   $\Delta H = -55.8 \text{ kJ}$  as shown note that the enthalpy change  $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$  is positive if heat is absorbed that is if  $\Delta H_{\text{products}}$

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web 3 12 hess s law in this experiment you will measure the amount of heat released in these three related exothermic reactions  $\text{NaOH}(\text{s}) + \text{Na}(\text{aq}) \rightarrow \text{OH}(\text{aq})$   $\text{H}_2 + 2\text{NaOH}(\text{s}) \rightarrow \text{H}_2(\text{aq}) + \text{Cl}(\text{aq})$   $\text{H}_2 + \text{O}(\text{Na}(\text{aq}) + \text{Cl}(\text{aq}) \rightarrow \text{H}_2 + 3$

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