

Trends on the Periodic Table

Part 1: Identifying chemical properties/characteristics of the elements

For each of the following elements, draw a caricature or cartoon above the description to show its "chemical personality" based on the description provided. (Note: See the Neon example for ideas.)

Lithium
Being one of the smaller elements with only 3 electrons, lithium is happy to give one of its electrons away and never seems to ask for any in return.



Francium
Francium is the largest of all of the elements and has so many electrons that it has trouble keeping track of them. Francium will eagerly give up electrons to other elements and never takes in extras. But watch out, it has an extremely reactive and explosive personality!

Fluorine
One of the smallest elements, Fluorine is often described as snappy with a "magnetic" personality. As part of the halogen family (outermost electron shell), Fluorine is the best at convincing other elements to give up their electrons to him. However, he holds on tight to his own electrons and will not let them go without a fight.



Iodine
As part of the halogen family, Iodine is also a skilled electron thief, however, not nearly as good as its smaller brothers and sisters. Iodine will also put up a pretty good fight to keep its own electrons.



Neon
Part of the Noble family, Neon does not readily interact with other elements. With a full set of valence electrons, Neon feels that he has everything he needs to make him happy without sharing with others.

Part 2: Use your textbook to complete the following:

Atomic Radii Size:

1. Explain **atomic radius**: A measure of the size (volume) of an atom. Determined by measuring half of the distance between two identical atoms bonded together.
2. Using the information in figure 5-13 (pg 141) in your book, **write** the atomic radii size for **alkali metals & period #2** on the periodic table to your right.



3. What happens to atomic radii as you move down a column and across a row? On the arrows on the diagram, indicate the trend. (Options: gets bigger OR gets smaller)
4. Of all the elements on the periodic table, **identify & circle the largest & smallest** in size.

5. Rank the following elements from largest to smallest in atomic radii:

a) Be, Ba, Mg = **Ba, Mg, Be**

b) Li, F, B = **Li, B, F**

6. How can you explain the trend?

a. Why do they get smaller as you go across a period? Additional protons in the nucleus "pull" on electrons, bring the electron cloud closer.

Effective nuclear charge, Z_{eff} , increases the attraction of the nucleus and therefore pulls the electron cloud closer to the nucleus resulting in smaller radii.

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4AP Answer

Periodic Trends Worksheet Answers Pogil

Thomas Griffiths



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Table of Contents Periodic Trends Worksheet Answers Pogil

1. Understanding the eBook Periodic Trends Worksheet Answers Pogil
 - The Rise of Digital Reading Periodic Trends Worksheet Answers Pogil
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trends Worksheet Answers Pogil
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Trends Worksheet Answers Pogil
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Trends Worksheet Answers Pogil
 - Personalized Recommendations
 - Periodic Trends Worksheet Answers Pogil User Reviews and Ratings
 - Periodic Trends Worksheet Answers Pogil and Bestseller Lists
5. Accessing Periodic Trends Worksheet Answers Pogil Free and Paid eBooks
 - Periodic Trends Worksheet Answers Pogil Public Domain eBooks
 - Periodic Trends Worksheet Answers Pogil eBook Subscription Services
 - Periodic Trends Worksheet Answers Pogil Budget-Friendly Options
6. Navigating Periodic Trends Worksheet Answers Pogil eBook Formats

- ePub, PDF, MOBI, and More
 - Periodic Trends Worksheet Answers Pogil Compatibility with Devices
 - Periodic Trends Worksheet Answers Pogil Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Trends Worksheet Answers Pogil
 - Highlighting and Note-Taking Periodic Trends Worksheet Answers Pogil
 - Interactive Elements Periodic Trends Worksheet Answers Pogil
 8. Staying Engaged with Periodic Trends Worksheet Answers Pogil
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Trends Worksheet Answers Pogil
 9. Balancing eBooks and Physical Books Periodic Trends Worksheet Answers Pogil
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Trends Worksheet Answers Pogil
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Periodic Trends Worksheet Answers Pogil
 - Setting Reading Goals Periodic Trends Worksheet Answers Pogil
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Periodic Trends Worksheet Answers Pogil
 - Fact-Checking eBook Content of Periodic Trends Worksheet Answers Pogil
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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