

Periodic Trends Worksheet

Directions: Use your notes to answer the following questions.

1. Rank the following elements by increasing atomic radius: carbon, aluminum, oxygen, potassium.

Oxygen < Carbon < Aluminum < Potassium

2. Rank the following elements by increasing electronegativity: sulfur, oxygen, neon, aluminum.

Neon < Aluminum < Sulfur < Oxygen

3. Why does fluorine have a higher ionization energy than iodine?

Fluorine has nine protons attracting 9 electrons which are much closer to the nucleus compared to iodine. This results in smaller size of Fluorine than iodine making it difficult to remove an electron from the atom.

4. Why do elements in the same family generally have similar properties?

Because they have same number of electrons in the outer shell (valence electrons) which take part in chemical reaction.

5. Indicate whether the following properties increase or decrease from left to right across the periodic table.

- | | |
|--|-----------|
| a. atomic radius (excluding noble gases) | Decreases |
| b. first ionization energy | Increases |
| c. electronegativity | Increases |

6. What trend in atomic radius occurs down a group on the periodic table? What causes this trend?

Atomic radius **decreases** down the group on the periodic table. As we compare the elements down the group, the effective nuclear charge increases, but at the same time the outermost electrons are found in the shell that is farther away from the nucleus. Also the number of screening electrons increases. This results in reduced attraction between the nucleus and the outermost electrons.

7. What trend in ionization energy occurs across a period on the periodic table? What causes this trend?

Ionization energy **increases** from left to right across the period. The atomic size becomes smaller from left to right. So it becomes harder to remove electron from the atom. Hence the energy required to do so (ionization Energy) increases.

Periodic Trends And Properties Of Elements Lab Answer Keys

ML Morrison



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Table of Contents Periodic Trends And Properties Of Elements Lab Answer Keys

1. Understanding the eBook Periodic Trends And Properties Of Elements Lab Answer Keys
 - The Rise of Digital Reading Periodic Trends And Properties Of Elements Lab Answer Keys
 - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Trends And Properties Of Elements Lab Answer Keys
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Periodic Trends And Properties Of Elements Lab Answer Keys
 - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Trends And Properties Of Elements Lab Answer Keys
 - Personalized Recommendations

- Periodic Trends And Properties Of Elements Lab Answer Keys User Reviews and Ratings
- Periodic Trends And Properties Of Elements Lab Answer Keys and Bestseller Lists
- 5. Accessing Periodic Trends And Properties Of Elements Lab Answer Keys Free and Paid eBooks
 - Periodic Trends And Properties Of Elements Lab Answer Keys Public Domain eBooks
 - Periodic Trends And Properties Of Elements Lab Answer Keys eBook Subscription Services
 - Periodic Trends And Properties Of Elements Lab Answer Keys Budget-Friendly Options
- 6. Navigating Periodic Trends And Properties Of Elements Lab Answer Keys eBook Formats
 - ePub, PDF, MOBI, and More
 - Periodic Trends And Properties Of Elements Lab Answer Keys Compatibility with Devices
 - Periodic Trends And Properties Of Elements Lab Answer Keys Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Periodic Trends And Properties Of Elements Lab Answer Keys
 - Highlighting and Note-Taking Periodic Trends And Properties Of Elements Lab Answer Keys
 - Interactive Elements Periodic Trends And Properties Of Elements Lab Answer Keys
- 8. Staying Engaged with Periodic Trends And Properties Of Elements Lab Answer Keys
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Periodic Trends And Properties Of Elements Lab Answer Keys
- 9. Balancing eBooks and Physical Books Periodic Trends And Properties Of Elements Lab Answer Keys
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Periodic Trends And Properties Of Elements Lab Answer Keys
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Periodic Trends And Properties Of Elements Lab Answer Keys
 - Setting Reading Goals Periodic Trends And Properties Of Elements Lab Answer Keys
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodic Trends And Properties Of Elements Lab Answer Keys
 - Fact-Checking eBook Content of Periodic Trends And Properties Of Elements Lab Answer Keys

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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