				Name	
		Worksheet:	: Periodic Trends	Period	
Which statement best describes Group 2 elements as they are considered in order from top to bottom of the			11. Which sequence of elements is arranged in order of decreasing atomic radii?		
	iodic Table?	er mour rep to content on the	(A) Al Si P	(C) Cl. Br. I	
(A) The number of principal energy levels increases, and		(B) Li, Na, K	(D) N. C. B		
the number of valence electrons increases.			(6) Lt. Na. K	(D) N. C. B	
(8)		pal energy levels increases, and	12. Which list of element	from Group 2 on the Periodic	
the number of valence electrons remains the same.			Table is arranged in order	of increasing atomic radius?	
(C) The number of principal energy levels remains the			(A) Be, Mg, Ca	(C) Ba, Ra, Sr	
same, and the number of valence electrons increases. (D) The number of principal energy levels remains the			(B) Ca, Mg, Be	(D) Sr. Ra. Ba	
19,000		r of valence electrons decreases.			
			13. As each successive element in Group 15 of the Periodic Table is considered in order of increasing atomic		
What is the total number of valence electrons in an			number, the atomic radius		
	m of boron in the grou		(A) decreases	(C) remains the same	
CA		(C) 3	(B) increases		
(19)	1.7	(D) 5			
100	What is the test of some	iber of valence electrons in an		m's attraction for the electrons in	
	atom of xenon, Xe?		a chemical bond is the atom's		
CAL	0	(C) 8	(A) electronegativity	(C) heat of reaction	
	2	(D) 18	(B) ionization energy	(D) heat of formation	
			15 Which proportion and	most common in nonmetals?	
4. The elements calcium and strontium have similar			(A) low ionization energy		
chemical properties because they both have the same			(B) low ionization energy and high electronegativity		
(A) atomic number				and low electronegativity	
	(B) mass number (C) number of valence electrons		(D) high ionization energy and high electronegativity		
(D) number of completely filled sublevels			16. Which Group 17 element has the least attraction for		
(12) manual or companies into subservers					
5. On the Periodic Table of the Elements, all the			electrons?	and the second second	
elie	ments within Group 16	have the same number of	(A) F	(C) Br	
ILA.	valence electrons	(C) protons	(B) Cl	(D) I	
(18)	energy levels	(D) neutrons	17. Which element in Gre	sup 16 has the greatest tendency	
- 2	An observe when a	rtially filled d sublevel in the	to gain electrons?		
	and state is classified a		(A) Te	(C) S	
	a halogen	(C) an alkali metal	(B) Sc	(D) O	
1000000	a transition metal	(D) an alkaline earth metal		n - Militaria - San	
88.65			18. The Group 17 elemen electronogativity is	t with the highest	
7. Which electron configuration represents a transition		(A) fluorine	(C) bromine		
	microf7		(B) chlorine	(D) iodine	
	1 x ² 2x ² 2p ²	(C) [Ar]3d ⁶ 4s ²	(B) chloring	(D) rodine	
(EK)	[Ne]3s ²	(D) [Ar]3d ¹⁰ 4s ² 4p ⁶	19. As the elements of Gr	oup 1 on the Periodic Table are	
-	Which element in Period 5 of the Periodic Table is a			considered in order of increasing atomic radius, the	
8. Which element in Period 5 of the Periodic Table is a transition element?		ionization energy of each successive element generally			
	Sr	(C) Ag	(A) decreases	(C) remains the same	
	Sb	(D) No	(B) increases		
		777	20. The amount of some	recognized to response the	
		ng atoms has the largest atomic		20. The amount of energy required to remove the outermost electron from a gaseous atom in the ground state	
	radius?		is known as		
(A	Na Na	(C) Mg	(A) first ionization energy	(C) conductivity	

(B) K.

(ID): Ca-

10. Which noble gas has the highest first ionization energy?

(A) radion

(C) meon

(B) krypton.

(D) helium

21. Which element is a member of the halogen family? (A) K.

(B) activation energy

(B) B

(C) I

(D) electronegativity

(D) S

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