Name _____ Date _____ Class _____

CHAPTER 5 REVIEW

The Periodic Law

SECTION 1

SHORT ANSWER **Answer the following questions in the space provided.**

1. ___c___ In the modern periodic table, elements are ordered

(a) according to decreasing atomic mass.
(b) according to Mendeleev's original design.
(c) according to increasing atomic number.
(d) based on when they were discovered.

2. ___d___ Mendeleev noticed that certain similarities in the chemical properties of elements appeared at regular intervals when the elements were arranged in order of increasing

(a) density.
(b) reactivity.
(c) atomic number.
(d) atomic mass.

3. ___b___ The modern periodic law states that

(a) no two electrons with the same spin can be found in the same place in an atom.
(b) the physical and chemical properties of an element are functions of its atomic number.
(c) electrons exhibit properties of both particles and waves.
(d) the chemical properties of elements can be grouped according to periodicity, but physical properties cannot.

4. ___c___ The discovery of the noble gases changed Mendeleev's periodic table by adding a new

(a) period.
(b) series.
(c) group.
(d) level.

5. ___d___ The most distinctive property of the noble gases is that they are

(a) metallic.
(b) radioactive.
(c) metalloid.
(d) largely unreactive.

6. ___c___ Lithium, the first element in Group 1, has an atomic number of 3. The second element in this group has an atomic number of

(a) 4.
(b) 10.
(c) 11.
(d) 18.

7. An isotope of fluorine has a mass number of 19 and an atomic number of 9.

___9___ **a.** How many protons are in this atom?

___10___ **b.** How many neutrons are in this atom?

___$^{19}_{9}F$___ **c.** What is the nuclear symbol of this fluorine atom, including its mass number and atomic number?

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Table of Contents Periodic Law Modern Chemistry Section Quiz Answers

1. Understanding the eBook Periodic Law Modern Chemistry Section Quiz Answers
    - The Rise of Digital Reading Periodic Law Modern Chemistry Section Quiz Answers
    - Advantages of eBooks Over Traditional Books
2. Identifying Periodic Law Modern Chemistry Section Quiz Answers
    - Exploring Different Genres
    - Considering Fiction vs. Non-Fiction
    - Determining Your Reading Goals
3. Choosing the Right eBook Platform
    - Popular eBook Platforms
    - Features to Look for in an Periodic Law Modern Chemistry Section Quiz Answers
    - User-Friendly Interface
4. Exploring eBook Recommendations from Periodic Law Modern Chemistry Section Quiz Answers
    - Personalized Recommendations
    - Periodic Law Modern Chemistry Section Quiz Answers User Reviews and Ratings
    - Periodic Law Modern Chemistry Section Quiz Answers and Bestseller Lists
5. Accessing Periodic Law Modern Chemistry Section Quiz Answers Free and Paid eBooks
    - Periodic Law Modern Chemistry Section Quiz Answers Public Domain eBooks
    - Periodic Law Modern Chemistry Section Quiz Answers eBook Subscription Services
    - Periodic Law Modern Chemistry Section Quiz Answers Budget-Friendly Options

6. Navigating Periodic Law Modern Chemistry Section Quiz Answers eBook Formats
    - ePub, PDF, MOBI, and More
    - Periodic Law Modern Chemistry Section Quiz Answers Compatibility with Devices
    - Periodic Law Modern Chemistry Section Quiz Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Periodic Law Modern Chemistry Section Quiz Answers
    - Highlighting and Note-Taking Periodic Law Modern Chemistry Section Quiz Answers
    - Interactive Elements Periodic Law Modern Chemistry Section Quiz Answers
8. Staying Engaged with Periodic Law Modern Chemistry Section Quiz Answers
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Periodic Law Modern Chemistry Section Quiz Answers
9. Balancing eBooks and Physical Books Periodic Law Modern Chemistry Section Quiz Answers
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Periodic Law Modern Chemistry Section Quiz Answers
10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
11. Cultivating a Reading Routine Periodic Law Modern Chemistry Section Quiz Answers
    - Setting Reading Goals Periodic Law Modern Chemistry Section Quiz Answers
    - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Periodic Law Modern Chemistry Section Quiz Answers
    - Fact-Checking eBook Content of Periodic Law Modern Chemistry Section Quiz Answers
    - Distinguishing Credible Sources
13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
14. Embracing eBook Trends
    - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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