

Periodic Trends Worksheet

Directions: Use your notes to answer the following questions.

1. Rank the following elements by increasing atomic radius: carbon, aluminum, oxygen, potassium.

Oxygen < Carbon < Aluminum < Potassium

2. Rank the following elements by increasing electronegativity: sulfur, oxygen, neon, aluminum.

Neon < Aluminum < Sulfur < Oxygen

3. Why does fluorine have a higher ionization energy than iodine?

Fluorine has nine protons attracting 9 electrons which are much closer to the nucleus compared to iodine. This results in smaller size of Fluorine than iodine making it difficult to remove an electron from the atom.

4. Why do elements in the same family generally have similar properties?

Because they have same number of electrons in the outer shell (valence electrons) which take part in chemical reaction.

5. Indicate whether the following properties increase or decrease from left to right across the periodic table.

- | | |
|--|------------------|
| a. atomic radius (excluding noble gases) | Decreases |
| b. first ionization energy | Increases |
| c. electronegativity | Increases |

6. What trend in atomic radius occurs down a group on the periodic table? What causes this trend?

Atomic radius **increases** down the group on the periodic table. As we compare the elements down the group, the effective nuclear charge increases, but at the same time the outermost electrons are found in the shell that is farther away from the nucleus. Also the number of screening electrons increases. This results in reduced attraction between the nucleus and the outermost electrons.

7. What trend in ionization energy occurs across a period on the periodic table? What causes this trend?

Ionization energy **increases** from left to right across the period. The atomic size becomes smaller from left to right. So it becomes harder to remove electron from the atom. Hence the energy required to do so (ionization Energy) increases.

Periodic Lab Table Trend Activities Answer Key

Hussin A.Rothana



Periodic Lab Table Trend Activities Answer Key:

The book delves into Periodic Lab Table Trend Activities Answer Key. Periodic Lab Table Trend Activities Answer Key is a crucial topic that must be grasped by everyone, from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into Periodic Lab Table Trend Activities Answer Key, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:
 - Chapter 1: Introduction to Periodic Lab Table Trend Activities Answer Key
 - Chapter 2: Essential Elements of Periodic Lab Table Trend Activities Answer Key
 - Chapter 3: Periodic Lab Table Trend Activities Answer Key in Everyday Life
 - Chapter 4: Periodic Lab Table Trend Activities Answer Key in Specific Contexts
 - Chapter 5: Conclusion
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 3. In chapter 2, this book will delve into the foundational concepts of Periodic Lab Table Trend Activities Answer Key. This chapter will elucidate the essential principles that need to be understood to grasp Periodic Lab Table Trend Activities Answer Key in its entirety.
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 6. In chapter 5, the author will draw a conclusion about Periodic Lab Table Trend Activities Answer Key. This chapter will summarize the key points that have been discussed throughout the book.
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