

Periodic Trends Worksheet

Directions: Use the table to determine the following properties:

1. Write the following elements by increasing atomic radius: sodium, sulfur, aluminum, oxygen, potassium.

K, Na, Al, S, O

2. Write the following elements by increasing electronegativity: sodium, oxygen, sulfur, aluminum.

O, S, Al, Na

3. Which atom has the higher first ionization energy: Na or K?

Na has a higher first ionization energy than K because it is smaller.

4. Which two elements in the same family generally have similar properties?

There have been two cases each where a third element acts as a bridge between two groups.

5. Evaluate whether the following properties increase or decrease from left to right across the periodic table:

- a. atomic radius (increases to the right) **decreases**
- b. first ionization energy **increases**
- c. electronegativity **increases**

6. Which period in the periodic table shows atoms with the greatest atomic radius? Which element has the greatest?

Period 7, francium, cesium, barium, lanthanum, actinium, and the lanthanide series.

7. Which period in the periodic table shows atoms with the smallest atomic radius? Which element has the smallest?

Period 2, fluorine, neon, oxygen, nitrogen, carbon, boron, and beryllium.

8. Write the atoms in each group that have the largest atomic radius.

Li	Be
Na	Mg
K	Ca
Rb	Sr
Cs	Ba
Fr	Ra

Al	Si	P	S
Ge	As	Se	Br
Sb	Te	I	Xe

9. Write the atoms in each group that have the smallest atomic radius.

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Xiaolong Qi



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Table of Contents Physical Science If8767 Pg 63 Key

1. Understanding the eBook Physical Science If8767 Pg 63 Key
 - The Rise of Digital Reading Physical Science If8767 Pg 63 Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science If8767 Pg 63 Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science If8767 Pg 63 Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Physical Science If8767 Pg 63 Key
 - Personalized Recommendations
 - Physical Science If8767 Pg 63 Key User Reviews and Ratings
 - Physical Science If8767 Pg 63 Key and Bestseller Lists
5. Accessing Physical Science If8767 Pg 63 Key Free and Paid eBooks

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 - Physical Science If8767 Pg 63 Key eBook Subscription Services
 - Physical Science If8767 Pg 63 Key Budget-Friendly Options
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- ePub, PDF, MOBI, and More
 - Physical Science If8767 Pg 63 Key Compatibility with Devices
 - Physical Science If8767 Pg 63 Key Enhanced eBook Features
7. Enhancing Your Reading Experience
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 - Interactive Elements Physical Science If8767 Pg 63 Key
8. Staying Engaged with Physical Science If8767 Pg 63 Key
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science If8767 Pg 63 Key
9. Balancing eBooks and Physical Books Physical Science If8767 Pg 63 Key
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science If8767 Pg 63 Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Physical Science If8767 Pg 63 Key
- Setting Reading Goals Physical Science If8767 Pg 63 Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Physical Science If8767 Pg 63 Key
- Fact-Checking eBook Content of Physical Science If8767 Pg 63 Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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