

8 • Electron Configuration & the Periodic Table**AP QUESTION**

- (a) Write the ground state electron configuration (short form or long form) for an arsenic atom.

 $\text{As } 1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^3$ or $\text{As } [\text{Ar}] 4s^2 3d^{10} 4p^3$ or $[\text{Kr}] 3d^{10} 4s^2 4p^3$

- (b) Give one permissible set of four quantum numbers for each of the outermost electrons in a single As atom when it is in its ground state.

$$\begin{array}{|c|c|c|c|} \hline 4 & 1 & -1 & +\frac{1}{2} \\ \hline \end{array}$$

$$\begin{array}{|c|c|c|c|} \hline 4 & 1 & 0 & +\frac{1}{2} \\ \hline \end{array}$$

$$\begin{array}{|c|c|c|c|} \hline 4 & 1 & +1 & +\frac{1}{2} \\ \hline \end{array}$$
 $\left(\begin{array}{c} +\frac{1}{2} \\ -\frac{1}{2} \end{array} \right)$

$$\begin{array}{|c|c|c|c|} \hline 4 & 0 & 0 & +\frac{1}{2} \\ \hline \end{array}$$

$$\begin{array}{|c|c|c|c|} \hline 4 & 0 & 0 & -\frac{1}{2} \\ \hline \end{array}$$

- (c) Is an isolated arsenic atom in the ground state
- paramagnetic
- or diamagnetic? Explain briefly.

Paramagnetic $4p^3 = \uparrow \uparrow \uparrow$ half-filled orbitals = paramagnetic.

- (d) What ion would As be expected to form?
- As^{3-}
-
- Write the electron configuration for the arsenic ion.

 $\text{As } [\text{Ar}] 4s^2 3d^{10} 4p^6$ or $\text{As } [\text{Kr}] 3d^{10} 4s^2 4p^6$ As gains electrons and has the same electron configuration as Kr.

3-	2-	1-	He
N	O	F	Ne
P	S	Cl	Ar
As	Se	Br	Kr

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