

1. As you move from left to right on the periodic table, why does the atomic radius decrease?

- ☐ As more protons are added, the more tightly the surrounding electrons are drawn inwards.
- ☐ As more neutrons are added, the more tightly the surrounding electrons are drawn inwards.
- ☐ As more electrons are added, the more tightly they draw themselves towards the nucleus.
- ☐ As the mass of the nucleus rises, the more tightly the surrounding electrons are drawn inwards.

- ☐ It decreases.
- ☐ It decreases, then increases.
- ☐ It remains constant.
- ☐ It increases, then decreases.
- ☐ It increases.

The periodic table displays elements from Hydrogen (H) to Oganesson (Og). Each element is represented by its symbol, atomic number, and name. The table is organized into groups and periods, with colors used to categorize elements: red for alkali metals, orange for alkaline earth metals, yellow for transition metals, light green for post-transition metals, green for metalloids, blue for nonmetals, and purple for noble gases.

- ☐ Argon (Ar)
- ☐ Phosphorus (P)
- ☐ Sodium (Na)
- ☐ Aluminum (Al)
- ☐ Magnesium (Mg)

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