

## Chapter 5 The Periodic Table

## Section 5.2 The Modern Periodic Table

(pages 130–138)

This section explains the organization of the modern periodic table and discusses the general properties of metals, nonmetals, and metalloids.

## Reading Strategy (page 130)

**Previewing** Before you read, complete the table by writing two questions about the periodic table on pages 132–133. As you read, write answers to your questions. For more information on this reading strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.

| Questions About the Periodic Table                                        |                                                                                                                              |
|---------------------------------------------------------------------------|------------------------------------------------------------------------------------------------------------------------------|
| Question                                                                  | Answer                                                                                                                       |
| Students might ask:<br>What does atomic mass mean?                        | Atomic mass is a value that depends on the distribution of an element's isotopes in nature and the masses of those isotopes. |
| Why are two columns of elements placed before the main body of the table? | Two kinds of elements are placed before the table to make the table more compact.                                            |

Other possible questions: Why are there two numbering systems for the columns? Why is period 7 incomplete?

## The Periodic Law (pages 131–133)

1. Is the following sentence true or false? In the modern periodic table, elements are arranged by increasing number of protons.  
\_\_\_\_\_ **true** \_\_\_\_\_
2. Explain why the number of elements per period varies. \_\_\_\_\_  
The number of elements per period varies because the number of available orbitals increases from energy level to energy level.
3. Properties of elements repeat in a predictable way when atomic numbers are used to arrange elements into groups. This pattern of repeating properties is called the \_\_\_\_\_  
periodic law.

## Atomic Mass (page 134)

4. Label the four types of information supplied for chlorine in the diagram.

|    |          |
|----|----------|
| a. | 17       |
| b. | Cl       |
| c. | Chlorine |
| d. | 35.453   |

- Atomic number
- Element symbol
- Element name
- Atomic mass

# Periodic Table Section 5 2 Answer

**Tao Wei**



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