

Annotated periodic table of the elements

1 H1.008Hydrogen																		13 B10.81Boron	14 C12.01Carbon	15 N14.01Nitrogen	16 O16.00Oxygen	17 F18.99Fluorine	18 Ne20.18Neon												
3 Li6.94Lithium		4 Be9.012Beryllium																	19 Al26.98Aluminium	20 Si28.09Silicon	21 P30.97Phosphorus	22 S32.06Sulfur	23 Cl35.45Chlorine	24 Ar39.95Argon											
11 Na22.99Sodium		12 Mg24.31Magnesium																	25 Sc44.96Scandium	26 Ti47.87Titanium	27 V50.94Vanadium	28 Cr52.00Chromium	29 Mn54.94Manganese	30 Fe55.85Iron	31 Co58.93Cobalt	32 Ni58.69Nickel	33 Cu63.55Copper	34 Zn65.38Zinc	35 Ga69.72Gallium	36 Ge72.63Germanium	37 As74.92Arsenic	38 Se78.96Selenium	39 Br79.90Bromine	40 Kr83.80Krypton	
19 K39.10Potassium		20 Ca40.08Calcium																	39 Y88.91Yttrium	40 Zr91.22Zirconium	41 Nb92.91Niobium	42 Mo95.94Molybdenum	43 Tc98.91Technetium	44 Ru101.07Ruthenium	45 Rh101.07Rhodium	46 Pd106.32Palladium	47 Ag107.87Silver	48 Cd112.41Cadmium	49 In114.82Indium	50 Sn118.71Tin	51 Sb121.76Antimony	52 Te127.60Tellurium	53 I126.91Iodine	54 Xe131.29Xenon	
37 Rb85.47Rubidium		38 Sr87.62Strontium																	57 to 70	71 Lu174.97Lutetium	72 Hf178.49Hafnium	73 Ta180.95Tantalum	74 W183.84Tungsten	75 Re186.21Rhenium	76 Os190.23Osmium	77 Ir192.22Iridium	78 Pt195.08Platinum	79 Au196.97Gold	80 Hg200.59Mercury	81 Tl204.38Thallium	82 Pb207.2Lead	83 Bi208.98Bismuth	84 Po209Polonium	85 At210Astatine	86 Rn222Radon
55 Cs132.91Cesium		56 Ba137.33Barium																	89 to 102	103 La138.91Lanthanum	104 Rf181Rutherfordium	105 Db185Dubnium	106 Sg271Seaborgium	107 Bh270Bohrium	108 Hs277Hassium	109 Mt276Meitnerium	110 Ds285Darmstadtium	111 Rg281Roentgenium	112 Cn285Copernicium	113 Nh284Nihonium	114 Fl289Flerovium	115 Mc288Moscovium	116 Lv293Livermorium	117 Ts294Tennessine	118 Og294Oganesson

Group names*

- Hydrogen & the alkali metals
- Alkaline earth metals
- Coinage metals
- Volatile metals
- Isotopes
- Chalcogens
- Postcogens
- Chalcogens
- Halogens
- Nitrogen group

*Groups 3 to 10 are named after their first members i.e. Group 3 is the Scandium group

Main group
oxidation states
electronegativity
Mendeleev

Atomic number
and symbol
Atomic weight
and name

CLASS COUNTERPARTS

Metals	Nonmetals
Pre-transition	
Active	Active
Transition	Intermediate
Post-transition	Metals
Metals (Ru-Pt, Os-Au)	Metals

Active metals

- Pre-transition metals**
Light, reactive, mostly of low mechanical strength, melting and boiling points. The group 1 members are easily cut with a knife. Only Be, Mg and Al have any structural uses. Most form basic oxides (Be and Al are amphoteric).
- Rare earths**
Closely resemble Ca, Sr and Ba, but many are heavier and they form mostly pale coloured compounds. Similar to one another, esp. La to Lu (aka the lanthanides), and hard to separate for that reason. Abundant but widely dispersed; commercially viable concentrations are rare. Oxides are strongly to moderately basic. Sc, Yb and Lu exhibit weak amphotericism.
- Actinides**
Soft, dense and reactive metals. Those from Ra onwards are synthetic. Early actinides (Th-Am) show some similarities to transition metals and have basic, amphoteric, or acidic oxides; late actinides are more like lanthanides. Many actinides form coloured compounds.

Transition metals

- Moderate to high density metals, with high melting and boiling points, many have high hardness and mechanical strength. Those from Ir onwards are synthetic. Chemically show variable valency and strong tendency to form coordination and colored compounds. Oxides are basic, amphoteric or acidic, depending on oxidation state.
- Noble metals**
Resistant to corrosion and oxidation, unlike base metals such as Fe, Ni, Cu and Zn. Commonly found together in native form. Good catalytic properties. Very easily electroplated; insoluble oxides.
- Frontier metals**
Soft (or brittle), low strength, and with melting points lower than transition metals. Crystalline structures tend to show directional bonding, with generally greater complexity or lower nearest neighbours than other metals. All used or tried as soft metal lubricants. Most are chemically weak; a minority are moderately electropositive. Characterised, to varying degrees, by covalent bonding tendencies, amphotericism, and anionic species such as zincates, stannates, and bismuthates.

Metals

- Metalloids**
Metallic looking brittle, fair electrical conductors. All known in semiconducting forms. Moderate ionisation energies (750–1,000 kJ/mol), and electronegativity values (1.8–2.3). Mostly behave chemically as non-metals. Can form alloys with metals. Amphoteric or weakly acidic oxides.
- Intermediate nonmetals**
H and N are colorless gases; C, P, and Se are metalloids looking solids while S is yellow. The latter four are brittle but all known in malleable, plastic or ductile forms. Moderate to high ionisation energies, low to high electron affinities, moderate to high electronegativity values, relatively poor to moderately strong oxidising agents, tend to form covalent or metallic to near-metallic compounds with metals. Their oxides are acidic, or neutral (H₂O, CO, NO, N₂O).
- Group 1**
The alkali metals (Li, Na, K, Rb, Cs, Fr) and the alkaline earth metals (Be, Mg, Ca, Sr, Ba, Ra) are the most reactive metals in the periodic table. They are soft, shiny, and have low melting points. They react with water to form hydroxides and hydrogen gas. They are also highly reactive with oxygen and halogens.

Nonmetals

- Chalcogens**
O, S, Se, and Te are, respectively, colorless, pale yellow, and pale yellow-green gases. Br is a red-brown liquid. I is a metallic looking solid. Chemically tend to have high ionisation energies, electron affinities, and electronegativity values, and be relatively strong oxidising agents. Manifestations include oxygen's major association with the ubiquitous processes of corrosion and combustion, and the intrinsically corrosive nature of the non-metallic halogens. All five tend to form predominantly ionic compounds with metals. Their oxides are acidic.
- Noble gases**
Colorless, odorless, non-flammable gases with very low chemical reactivity. The first noble gas compound, of approximate composition XePtF₆, was not prepared until 1962. More than half-a-thousand noble gas compounds were synthesised in the ensuing half-century. Compounds of He, Ne, and Ar, in ambient conditions, are not known.

Radioactive
Oganesson is likely to be either a metalloid or a post-transition metal

Periodic Table Guide Answer Sheet

Michael Rajnik

A red circular graphic with a gradient, appearing as a partial circle or a thick arc, located to the right of the author's name.

Periodic Table Guide Answer Sheet:

Adopting the Beat of Phrase: An Psychological Symphony within **Periodic Table Guide Answer Sheet**

In some sort of used by screens and the ceaseless chatter of fast interaction, the melodic splendor and psychological symphony produced by the prepared term often diminish into the background, eclipsed by the relentless noise and interruptions that permeate our lives. However, located within the pages of **Periodic Table Guide Answer Sheet** a charming fictional value full of fresh feelings, lies an immersive symphony waiting to be embraced. Constructed by an elegant musician of language, this fascinating masterpiece conducts viewers on a psychological journey, well unraveling the hidden songs and profound impact resonating within each carefully constructed phrase. Within the depths of this moving review, we shall discover the book is key harmonies, analyze its enthralling writing style, and submit ourselves to the profound resonance that echoes in the depths of readers souls.

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