

Full Name _____ Date _____ Per _____ Chemistry Chapter 6 The Periodic Table

Periodic Trends Practice

1. Identify each element as a metal, metalloid, or nonmetal

germanium _____ metalloid _____

zinc _____ metal _____

phosphorous _____ nonmetal _____

lithium _____ metal _____

2. Give two examples of elements for each category

noble gases _____ He, Ne, Ar, Kr, Xe, Rn _____

halogens _____ F, Cl, Br, I, At _____

alkali metals _____ Li, Na, K, Rb, Cs, Fr _____

alkaline earth metals _____ Be, Mg, Ca, Sr, Ba, Ra _____

transition metals _____ Any elements from 3B-28 _____

3. What trend in atomic radius do you see as you go down a group/family on the periodic table? What causes this trend? Increases going down due to added energy levels. The inside energy levels are filled up so all remaining electrons

must go into open spaces further away from the nucleus.

4. What trend in atomic radius do you see as you go across a period/row on the periodic table? Decreases going across

due to electrons being added into the same valence shell. Protons would be added to the nucleus, making the positive charge stronger and electrons placed in the same space would be pulled in closer by the nucleus.

5. Circle the atom in each pair that has the largest atomic radius.

A. Al B

B. S O

C. Br Cl

D. Na Al

E. O F

F. Mg Ca

6. Rank the following elements by increasing atomic radius: carbon, aluminum, oxygen, potassium.

O < C < Al < K

7. Define ionization energy. The amount of energy required to remove an electron from a neutral atom

8. Is it easier to form a positive ion with an element that has a high ionization energy or an element that has a low ionization energy? Explain. It would be easier to form a + ion with a low ionization energy because less energy would need to be used to remove the electron.

9. What trend in ionization energy do you see as you go down a group/family on the periodic table? Going down a group the ionization energy decreases. The valence electrons are further away and easier to take away.

10. What trend in ionization energy do you see as you go across a period/row on the periodic table? What causes this trend? Going across the period, ionization energy increases. The atoms are smaller, so the nucleus is pulling electrons in closer making them harder to remove. The atoms are also increasing their number of valence electrons. As they get closer to having a complete set of 8 valence electrons they do not want to lose electrons any more, they want to gain them instead.

11. Circle the atom in each pair that has the greater ionization energy

A. Li Be

B. Ca Ba

C. Na K

D. P Ar

E. Cl Si

F. Li K

12. Define electronegativity. The ability of an atom to attract electrons in a bond.

13. What trend in electronegativity do you see as you go down a group/family on the periodic table? Values decrease within a group. Since the valence shell becomes further away, it is harder for the nucleus to draw electrons towards it.

14. What trend in electronegativity do you see as you go across a period/row on the periodic table? What causes this trend? Values increase across a period because the atoms are getting smaller and the nucleus can pull electrons in easier and the atoms are closer to being like noble gases. They want to be able to complete their valence shell.

Periodic Trends Summary Table Answer Key

Leo J. Malone, Theodore Dolter



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via your online Student Tools Premium Practice for AP Excellence 7 full length practice tests 5 in the book 2 online with detailed answer explanations Practice drills at the end of each content review chapter Review of important laboratory procedures and equipment *Prentice Hall Chemistry*, 2000 **Basic Concepts of Chemistry** Leo J. Malone, Theodore Dolter, 2008-12-03 Engineers who need to have a better understanding of chemistry will benefit from this accessible book It places a stronger emphasis on outcomes assessment which is the driving force for many of the new features Each section focuses on the development and assessment of one or two specific objectives Within each section a specific objective is included an anticipatory set to orient the reader content discussion from established authors and guided practice problems for relevant objectives These features are followed by a set of independent practice problems The expanded Making it Real feature showcases topics of current interest relating to the subject at hand such as chemical forensics and more medical related topics Numerous worked examples in the text now include Analysis and Synthesis sections which allow engineers to explore concepts in greater depth and discuss outside relevance *A Look at Zambia's Road Financing Strategies: The Impact of Incongruent Policies and Budgets* Yohane Tembo, 2015-01-01 There are various reasons why the road asset in Zambia has been deteriorating over the past five decades since independence Much effort to restore the road asset to its original value and keep it maintained in a good and safe condition was initiated in mid 1990s with the launch of the Road Sector Investment Plan Phase I which was immediately followed by the launch of a Phase II to run from 2003 to 2013 It is my appreciation that many diagnostic studies into road financing strategies which could have led to inefficient road asset management in Zambia have been undertaken and various solutions tendered The author takes cognizance that resource mobilization for road construction and maintenance and subsequent allocation to respective road programs can often be problematic for an emerging country like Zambia It has the potential to lead to inefficiencies in road asset management as has evidently been demonstrated in the declining road asset value over time The political as well as the economic landscape plays a vital role in resource mobilization and allocation strategies as much as the institutional and the legal framework do The failure to clear the backlog of maintenance which normally results largely from deferment of scheduled maintenance due to insufficient annual budgetary allocation to the road sector maintenance programs has led to the significant deterioration in road network condition Exorbitant road construction costs have posed additional challenges to the fiscus thereby constraining both the quantity and quality of road infrastructure that could be constructed and maintained at any given time Recent policy drives have been categorical in their preference of new road construction aimed at linking Zambia over road maintenance which plays a pivotal role in road asset management This creates a perception that policy pronouncements are at variance to policy documents which promote sustainable economic development through efficient road investments and could be seen as being paradoxical in that the actual financing strategies are skewed towards road construction a recipe of comfort for the next election challenge due to increased visibility on the ground rather than to maintenance as an astute and

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