

Chapter 3: Periodic Trends (NB p.29-30)

Vocabulary: Period, group, atomic radius, ionization energy.

Equations & Constants: N/A

Practice Problems:

- D 1. The energy required to remove an electron from an atom is the atom's
 a. electron affinity.
 b. electron energy.
 c. electronegativity.
 d. ☒ ionization energy.

- D 2. The element that has the greatest electronegativity is
 a. oxygen.
 b. sodium.
 c. chlorine.
 d. ☒ fluorine.

- A 3. One-half the distance between the nuclei of identical atoms that are bonded together is called the
 a. ☒ atomic radius.
 b. atomic diameter.
 c. atomic volume.
 d. electron cloud.

- C 4. A positive ion is known as a(n)
 a. ionic radius.
 b. valence electron.
 c. ☒ cation.
 d. anion.

- D 5. A negative ion is known as a(n)
 a. ionic radius.
 b. valence electron.
 c. cation.
 d. ☒ anion.

- A 6. Across a period in the periodic table, atomic radii
 a. ☒ gradually decrease.
 b. gradually decrease, then sharply increase.
 c. gradually increase.
 d. gradually increase, then sharply decrease.

- C 7. Which of the following elements has the largest atomic radius?
 a. Fluorine
 b. Bromine
 c. ☒ Iodine
 d. Chlorine

- A 8. Which of the following elements has the largest atomic radius?
 a. ☒ Chromium
 b. Manganese
 c. Zinc
 d. Iron

- C 9. Which of the following elements has the largest atomic radius?
 a. Calcium
 b. Magnesium
 c. ☒ Barium
 d. Strontium

- D 10. Which of the following elements has the largest ionization energy?
 a. Hydrogen
 b. Helium
 c. Lithium
 d. ☒ Beryllium

- A 11. Which of the following elements has the largest ionization energy?
 a. ☒ Phosphorous
 b. Arsenic
 c. Selenium
 d. Carbon

- B 12. Which of the following elements has the largest ionization energy?
 a. Calcium
 b. ☒ Magnesium
 c. Barium
 d. Strontium

Atomic Radius

1. Atoms get **larger** going down a group because adding new shells/energy levels

2. Atoms get **smaller** going across a period because not adding energy levels but $\uparrow Z_{eff}$

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Table of Contents Ofl Packet Answers English

1. Understanding the eBook Ofl Packet Answers English
 - The Rise of Digital Reading Ofl Packet Answers English
 - Advantages of eBooks Over Traditional Books
2. Identifying Ofl Packet Answers English
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ofl Packet Answers English
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ofl Packet Answers English
 - Personalized Recommendations
 - Ofl Packet Answers English User Reviews and Ratings
 - Ofl Packet Answers English and Bestseller Lists
5. Accessing Ofl Packet Answers English Free and Paid eBooks

- Ofl Packet Answers English Public Domain eBooks
- Ofl Packet Answers English eBook Subscription Services
- Ofl Packet Answers English Budget-Friendly Options
- 6. Navigating Ofl Packet Answers English eBook Formats
 - ePub, PDF, MOBI, and More
 - Ofl Packet Answers English Compatibility with Devices
 - Ofl Packet Answers English Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ofl Packet Answers English
 - Highlighting and Note-Taking Ofl Packet Answers English
 - Interactive Elements Ofl Packet Answers English
- 8. Staying Engaged with Ofl Packet Answers English
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ofl Packet Answers English
- 9. Balancing eBooks and Physical Books Ofl Packet Answers English
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ofl Packet Answers English
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Ofl Packet Answers English
 - Setting Reading Goals Ofl Packet Answers English
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Ofl Packet Answers English
 - Fact-Checking eBook Content of Ofl Packet Answers English
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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