

phosphorus(V) oxide with water

phosphoric (V) acid

P_4O_{10} → single molecular
→ van der Waals

MgO → giant ionic
→ ionic

[2]

(d) Explain why phosphorus(V) oxide has a low melting point of approximately 300°C but magnesium oxide has a high melting point of approximately 2850°C .

The melting of phosphorus(V) oxide involve the breaking of weak van der Waals' forces between molecules.

The melting of magnesium oxide involve the breaking of electrostatic forces between cation and anion which is ionic bonds.

The ionic bonds is stronger than vdw forces.

[3]

solved by TheamKL

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