

## CLO- Classification of Elements & Periodicity in Properties

### MULTIPLE CHOICE QUESTIONS:

Choose the one correct option in the following:

- Identify the least stable ion amongst the following :  
(a)  $\text{Li}^+$  (b)  $\text{Be}^+$  (c)  $\text{B}^+$  (d)  $\text{C}^+$
- The correct order of radii is :  
(a)  $\text{N} < \text{Be} < \text{B}$  (b)  $\text{F}^- < \text{O}^{2-} < \text{N}^{3-}$  (c)  $\text{N} < \text{Li} < \text{K}$  (d)  $\text{Fe}^{2+} < \text{Fe}^{3+} < \text{Fe}^{4+}$
- Which one of the following has largest size :  
(a)  $\text{Na}$  (b)  $\text{Na}^+$  (c)  $\text{Na}^-$  (d)  $\text{Li}^+$
- The number of elements in the 4<sup>th</sup> period of periodic table is :  
(a) 8 (b) 10 (c) 18 (d) 32
- Which of the following has least electron affinity (electron gain enthalpy) :  
(a)  $\text{O}$  (b)  $\text{N}$  (c)  $\text{Ar}$  (d)  $\text{F}$
- An element of atomic number 29 belongs to which of the following block of the periodic table :  
(a) s-block (b) p-block (c) d-block (d) f-block
- Which one of the following arrangements not truly represents the property indicated against it:  
(a)  $\text{Br}_2 < \text{Cl}_2 < \text{F}_2$  – electronegativity (b)  $\text{Br}_2 < \text{F}_2 < \text{Cl}_2$  – Electron Affinity  
(c)  $\text{Br}_2 < \text{Cl}_2 < \text{F}_2$  – Bond energy (d)  $\text{Br}_2 < \text{Cl}_2 < \text{F}_2$  – Oxidising Power
- Alkali metals in each period have :  
(a) Smallest size (b) Lowest IE (c) Highest IE (d) Highest EN
- A sudden large jump between the value of second and third ionization energies of elements would be associated with which of the following electronic configuration :  
(a)  $2\text{p}^4 3\text{s}^1$  (b)  $3\text{s}^2 3\text{p}^1$  (c)  $3\text{s}^2 3\text{p}^2$  (d)  $2\text{p}^4 3\text{s}^2$
- Which of the following pairs are chemically dissimilar :  
(a)  $\text{Na}$  and  $\text{K}$  (b)  $\text{Ba}$  and  $\text{Sr}$  (c)  $\text{Zr}$  and  $\text{Hf}$  (d)  $\text{Ca}$  and  $\text{Zn}$
- Collective name given to the elements with outer shell configuration  $ns^2 np^6$  is :  
(a) Chalcogens (b) Alkaline earths metals (c) Transition metals (d) Nobel gases
- Modern extended periodic table was given by :  
(a) Bohr (b) Lothar mayer (c) Moseley (d) None alone
- The most electronegative elements is :  
(a) Hydrogen (b) Fluorine (c) Chlorine (d) Oxygen
- The element having electronic configuration :  
 $[\text{Kr}]4\text{d}^{10}4\text{f}^{14}5\text{s}^2 5\text{p}^6 5\text{d}^1 6\text{s}^2$  belongs to  
(a) s-block (b) p-block (c) d-block (d) f-block
- An elements R forms the highest oxide  $\text{R}_2\text{O}_5$ . R belong to :

# Periodicity Multiple Choice Questions

**Michael Rajnik**



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