

CH 5.3 - Periodic Trends WebQuest PART I

There is a .pdf version of this assignment on my blog. Please utilize the online version to link the webpages in order to avoid typing in the addresses. Reminder: columns on the periodic table are called families or groups and rows are called periods

Atomic Radius (pages 174-175)

Using your text, please answer the following:

1. Can the radius of an atom be measured directly? _____
2. In what unit is atomic radius usually measured? _____
3. Define Atomic Radius: _____

Go to the following website:

http://www.webelements.com/periodicity/atomic_radius_empirical/bar_chart.html

4. Look at the graph. What is on the X Axis? _____
5. What is on the Y Axis (include units)? _____
6. **The colors on the graph represent rows on the periodic table.**
In a complete sentence, answer - What appears to be the trend in atomic radius as you move from left to right in a row?

In a complete sentence, answer - What appears to be the trend in atomic radius as you move down a column?

7. Using your text, please answer the following:

- a. What is the general trend for atomic size within a group/family from top to bottom?
- b. What two variables affect atomic size within a group?
- c. What is the general trend in atomic size from left to right within a period?

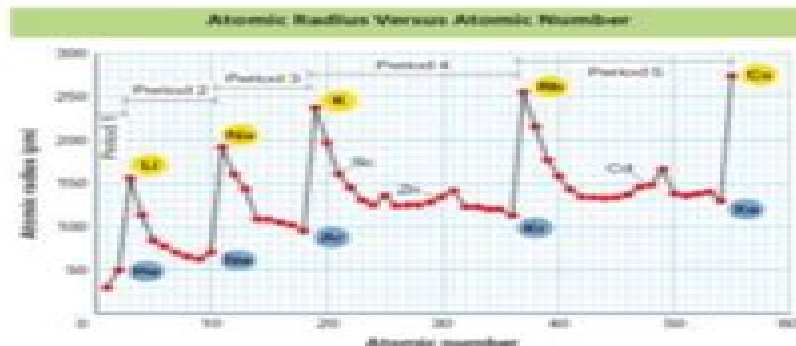
8. Use the graph to the right to answer the following:

- a. Determine which metal has an atomic radius of 238 pm?
- b. Is an atom of barium, atomic number 56, smaller or larger than an atom of cesium (Cs)?
- c. Which has a larger atom, helium or argon?
- d. Which has a larger atom, potassium or argon?

Teacher Checkpoint _____

Ions and Ionization Energy (pages 177-181)

Using your text, please answer the following:



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A Gutmann



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