

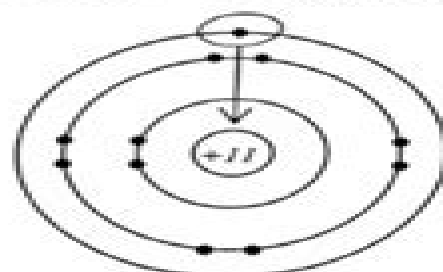
## Advanced Periodic Trends

How do the particles in an atom interact to dictate periodic properties of elements?

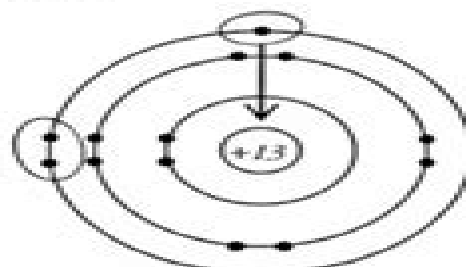
### Why?

Previously you have learned about Coulombic attraction and how it governs the trends in properties among the elements of the Periodic Table. For example, as you move across a period, atoms have more protons in the nucleus, which pulls the electrons in tighter making smaller, more electronegative atoms. As you proceed down a column of the periodic table, the valence electrons get further away, and the atoms get larger and less electronegative. Inquisitive students will always ask, "If more protons in the nucleus pull electrons in tighter, then why don't atoms get smaller when going down a column?" That is a good question, and one that you will investigate in this activity.

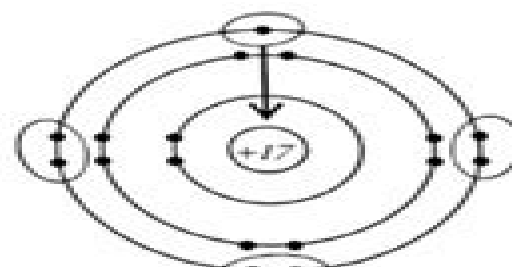
### Model 1 – Period 3 Elements



Sodium



Aluminum



Chlorine

|                             |            |            |             |
|-----------------------------|------------|------------|-------------|
| Atomic Number               | 11         | 13         | 17          |
| Number of Electrons         | 11         | 13         | 17          |
| Number of Valence Electrons | 1          | 3          | 7           |
| Core Charge                 | +1         | +3         | +7          |
| Atomic Radius               | 186 pm     | 143 pm     | 99 pm       |
| 1st Ionization Energy       | 496 kJ/mol | 578 kJ/mol | 1251 kJ/mol |
| Electronegativity           | 0.9        | 1.5        | 3.0         |

1. Describe the relationship of the three elements in Model 1 with regard to their relative positions on the periodic table.

*The three elements are in the same row of the periodic table.*

2. Refer to a periodic table to complete the first three rows of the table in Model 1.

*See Model 1.*

# Periodic Trends Pogil Pleasanton

**H Kauffman**



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