

THE PERIODIC TABLE OF ELEMENTS

TEACHER ANSWER KEY *vocabulary practice*

WORDS:

Period/row
reactive

group/family
nonreactive

metal
conduct

nonmetal
atomic mass

metalloid

PART I: MAGIC SQUARE

Directions: Match the number of the statement best describing the word in the box and write it on the line below the word. IF they are all right, each column and row will total the same MAGIC NUMBER. Good luck!

reactive • <u>8</u>	metal • <u>3</u>	metalloid • <u>6</u>	<u>15</u>
atomic mass • <u>1</u>	group/family • <u>5</u>	nonmetal • <u>7</u>	<u>15</u>
conduct • <u>4</u>	nonreactive • <u>9</u>	period/row • <u>2</u>	<u>15</u>

TOTAL: 15 15 15

- Typically shiny, hard, solid elements that easily conduct heat and electricity. The majority of elements on the periodic table are this type of element.
- There are 7 of these on the PTE. Here you can see patterns such as increasing atomic number from left to right and decreasing reactivity of elements from left to right.
- On an element square, this number tells the total number of protons + neutrons.
- In science with relation to the periodic table, this is the ability to generate or create heat and/or electricity.
- Elements in this share similar properties, meaning they react and behave similarly to other elements found in the same column.
- An element that looks like a metal, but is a poor conductor of electricity.
- Elements that do not contain or resemble characteristics of metal. They are poor conductors of electricity.
- An element that easily and readily responds (joining, forming new substances) to other elements.
- An element that does not respond noticeably to the presence of other elements.

OVER →

Periodic Table Vocabulary Answer Key

Chao Zhang

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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Hands-On General Science Activities With Real-Life Applications Pam Walker, Elaine Wood, 2008-04-21 In this second edition of Hands On General Science Activities with Real Life Applications Pam Walker and Elaine Wood have completely revised and updated their must have resource for science teachers of grades 5-12. The book offers a dynamic collection of classroom ready lessons, projects, and lab activities that encourage students to integrate basic science concepts and skills into everyday life.

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The Practice of Chemistry Donald J. Wink, Sharon Fetzer-Gislason, Sheila McNicholas, 2003-03. Students can't do chemistry if they can't do the math. The Practice of Chemistry First Edition is the only preparatory chemistry text to offer students targeted, consistent mathematical support to make sure they understand how to use math, especially algebra, in chemical problem solving. The book's unique focus on actual chemical practice, extensive study tools, and integrated media makes The Practice of Chemistry the most effective way to prepare students for the standard general chemistry course and bright futures as science majors.

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