

Quiz & Worksheet - Atomic and Ionic Radii on the Periodic Table

1. As you move from left to right on the periodic table, why does the atomic radius decrease?

- As more protons are added, the more tightly the surrounding electrons are drawn inwards.
- As more neutrons are added, the more tightly the surrounding electrons are drawn inwards.
- As more electrons are added, the more tightly they draw themselves towards the nucleus.
- As the mass of the nucleus rises, the more tightly the surrounding electrons are drawn inwards.

2. As you move from top to bottom on the periodic table, what happens to the atomic radius?

- in. It decreases.
- It decreases, then increases.
- tt nemains constant.
- It increases, then decreases.
- It increases.

Which of these elements has the largest atomic radius: argon, phosphorus, aluminum, sodium or magnesium?



- in: Argon (Art)
- Phosphorus (P)
- Seedium (Nati
- Altuminum (Alb
- Magnesium (Mg)

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Periodic Elements Need To Knows Answer Key

American Chemical Society

Periodic Elements Need To Knows Answer Key:

Decoding **Periodic Elements Need To Knows Answer Key**: Revealing the Captivating Potential of Verbal Expression

In a period characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its capability to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Periodic Elements Need To Knows Answer Key**," a mesmerizing literary creation penned by a celebrated wordsmith, readers attempt an enlightening odyssey, unraveling the intricate significance of language and its enduring effect on our lives. In this appraisal, we shall explore the book is central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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