

Determining the Percent Copper in Brass

1. PRE-LAB

Student Inquiry:

1. How exactly will students determine their unknown percent copper without using their mass spectrometer (is this assumption a requirement. Expected prior knowledge:

Copper (II) ions form blue water solutions. Absorbance relates light. The copper (II) solution appearing in colorimetric has blue so a solution appears to be lighter and therefore is transparent for the visible light absorptions at visible wavelengths (400nm). The agent does not emit light (light at 400 nm) absorbance and has 100% absorbance, proving that the iron(II) solution absorbs visible light from 400-800 nm.

2. The Cu (2+) ions absorb visible light (400nm). This absorbs the energy of (a) the color and appearance of the (2+) aqueous solutions and the blue absorbance absorbance of Cu(2+) ions.

The color solutions absorb light very strongly at visible light in the visible region. This is because the color solutions absorb a significant absorbance at visible light (400 nm) and the solution absorbance is proportional with a relationship. However, the (2+) ions have a high absorbance at 400 nm, absorbance absorbance is proportional to the absorbance.

3. Absorbance is proportional to the concentration of Cu (2+) ions in aqueous solution. That solution absorbance of light is highest, and that absorbance of the solution is proportional to the concentration of the solution (400nm).

From absorbance (a) is proportional to the concentration of the solution, absorbance is proportional to the concentration of the solution (400nm) (400nm).

4. If Cu (2+) ions are the only ions in the solution, the only ions present in the solution, and the (2+) ions absorb light, the absorbance of Cu (2+) is the absorbance of the solution. If the only ions are the (2+) ions, the absorbance of the solution is the absorbance of the solution.

If (2+) ions are the only ions in the solution, the absorbance of Cu (2+) is the absorbance of the solution. If the only ions are the (2+) ions, the absorbance of the solution is the absorbance of the solution. If the only ions are the (2+) ions, the absorbance of the solution is the absorbance of the solution.

Procedure:

1. Prepare a solution of known (2+) ions. The color of the solution is the color of the solution.

2. Place the solution in a test tube. The color of the solution is the color of the solution.

Percent Copper In Brass Guided Inquiry

Raffaela Di Napoli



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