

Percent, Actual, and Theoretical Yield

Example 1: 100g of CaCO_3 is heated.
 (a) Write the balanced equation for the reaction.
 (b) Calculate the theoretical yield of CaO if 100g of CaCO_3 is heated.



Example 2: 100g of CaCO_3 is heated.
 (a) If 56g of CaO is produced, what is the percent yield of CaO ?

$$\text{Percent Yield} = \frac{\text{Actual Yield}}{\text{Theoretical Yield}} \times 100\% = \frac{56 \text{ g CaO}}{56.08 \text{ g CaO}} \times 100\% = 99.86\%$$

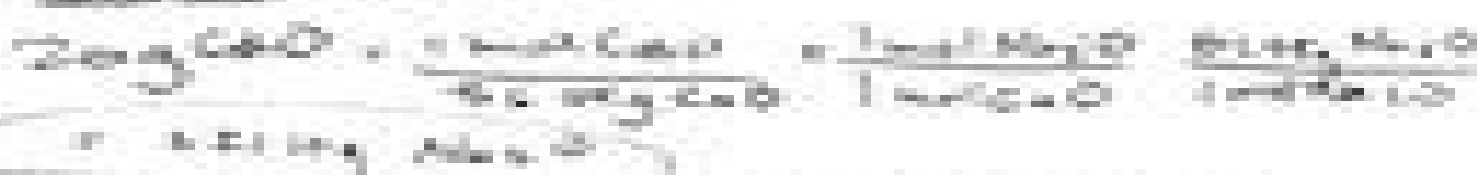
Example 3: 100g of CaCO_3 is heated.
 (a) Calculate the percent yield of CaO if 56g of CaO is produced.



Example 4: 100g of CaCO_3 is heated.
 (a) Calculate the percent yield of CaO if 56g of CaO is produced.

$$\text{Percent Yield} = \frac{\text{Actual Yield}}{\text{Theoretical Yield}} \times 100\% = \frac{56 \text{ g CaO}}{56.08 \text{ g CaO}} \times 100\% = 99.86\%$$

Example 5: 100g of CaCO_3 is heated.
 (a) Calculate the percent yield of CaO if 56g of CaO is produced.



Percent Actual And Theoretical Yield Answer Key

Ying-Ying Zheng



Percent Actual And Theoretical Yield Answer Key:

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