Model 2 – Ternary Acids (Oxyacids)

Acid	Name of Acid in Aqueous Solution	Cation (+)	Polyatomic Anion	Polyatomic Anion Name
HCIO.	Chloric acid	H,O	C/O3-1	Chlorate
H.SO,	Sulfurous acid	24,00	503-2	Sulfire
H.SO.	Sulfuric acid	2 Hso*	504-3	Sulfare
H,PO,	Phosphorous acid	3 H,O	P03-3	Phosphite
H.PO.	Phosphoric acid	3 H50+	PO _a s-	Phosphate
HNO,	Nitric acid	H30+	NO."	Nitrose
HNO,	Nitrous acid	H.0*	NO,	Nitrite
H,CO,	Carbonic acid	2450+	CO,-2	Carbonate

- 6. Look at the formulas of the ternary acids in Model 2.
 - w. How are ternary acids different from binary acids in their structure?

Ternory acids contain three elements and they also contain the element oxygen

6. What number do you think the prefix "ter-" refers to?

"ter" rofers to three

- 7. When ternary acids are mixed with water, ions will form. Fill in the table above with the formulas and names of the anions.
- 8. Examine the pairs of ternary acids in Model 2 that contain sulfur, phosphorus, and nitrogen. Each pair has one acid that ends in "-ic" and another that ends in "-ous." These endings are related to the name of the polyatomic anion found in the acid ("-ate" or "-ite"). Complete the statements below with the correct acid name ending.

Polyatomic anion ending is "-ate" → acid name ending is /C

Polyatomic anion ending is "-ite" → acid name ending is 045 ...

9. If the prefix "hydro-" were used to name a ternary acid, what problem would this create when

- naming HCIO? CIO3' has the name of Chlorate. "ate" must be replaced with "ic" If "hydro" is used as a prefix, it would be hydrochloric acid. This has the formula HCI, not HCIO3.

 The prefix "hydro" is not used in the name when an acid contain more than two elements and one of the elements is oxygen.

 "ate is replicate with it and "ite" is replicate with "ous".

Circle the acid(s) below that would be named beginning with the prefix "hydro-."

FESO.

H,CO.

HNO.

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RD Boyd

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- 1. Understanding the eBook Polyatomic Ions Worksheet Pogil
 - The Rise of Digital Reading Polyatomic Ions Worksheet Pogil
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Polyatomic Ions Worksheet Pogil
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Polyatomic Ions Worksheet Pogil
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Polyatomic Ions Worksheet Pogil
 - Personalized Recommendations
 - Polyatomic Ions Worksheet Pogil User Reviews and Ratings
 - Polyatomic Ions Worksheet Pogil and Bestseller Lists
- 5. Accessing Polyatomic Ions Worksheet Pogil Free and Paid eBooks
 - $\circ\,$ Polyatomic Ions Worksheet Pogil Public Domain eBooks
 - Polyatomic Ions Worksheet Pogil eBook Subscription Services
 - Polyatomic Ions Worksheet Pogil Budget-Friendly Options
- 6. Navigating Polyatomic Ions Worksheet Pogil eBook Formats

- o ePub, PDF, MOBI, and More
- Polyatomic Ions Worksheet Pogil Compatibility with Devices
- Polyatomic Ions Worksheet Pogil Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Polyatomic Ions Worksheet Pogil
 - Highlighting and Note-Taking Polyatomic Ions Worksheet Pogil
 - Interactive Elements Polyatomic Ions Worksheet Pogil
- 8. Staying Engaged with Polyatomic Ions Worksheet Pogil
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Polyatomic Ions Worksheet Pogil
- 9. Balancing eBooks and Physical Books Polyatomic Ions Worksheet Pogil
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Polyatomic Ions Worksheet Pogil
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Polyatomic Ions Worksheet Pogil
 - Setting Reading Goals Polyatomic Ions Worksheet Pogil
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Polyatomic Ions Worksheet Pogil
 - Fact-Checking eBook Content of Polyatomic Ions Worksheet Pogil
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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