

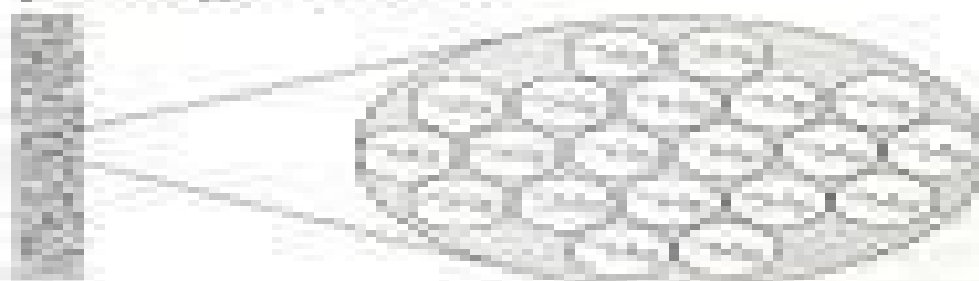
Average Atomic Mass

How can the masses of the particles tell us something?

Why?

Many different types of atoms show up naturally in our lives around us. For example, hydrogen, the lightest of them, is everywhere. Even the water around us is made of hydrogen. Hydrogen is part of almost everything that is living in this world. Atoms of the same element have different masses. For example, carbon has two different isotopes. One isotope of carbon has 6 protons and 6 neutrons, while the other has 6 protons and 7 neutrons. The mass of the first isotope is 12 amu, and the mass of the second is 13 amu. The average atomic mass of carbon is 12.01 amu. This is because the first isotope is more abundant than the second.

Model 1 - A Strip of Magnesium Metal



1. What are the masses of the two isotopes of magnesium shown in Model 1?
They are 24 and 26 amu.
2. What are the masses of the two isotopes of magnesium shown in Model 1?
24 and 26 amu.
3. How many of each isotope of magnesium are shown in Model 1? How many protons and neutrons are in each?
There are 7 Mg-24 atoms and 3 Mg-26 atoms. Each Mg-24 atom has 12 protons and 12 neutrons. Each Mg-26 atom has 12 protons and 14 neutrons.
4. How many protons and neutrons are in each of the two isotopes of magnesium shown in Model 1? How many protons and neutrons are in each?
Each Mg-24 atom has 12 protons and 12 neutrons. Each Mg-26 atom has 12 protons and 14 neutrons.

	Mg-24	Mg-26	Mg-24
Number of atoms	7	3	10

5. What is the mass of the two isotopes of magnesium shown in Model 1?
24 and 26 amu.
 6. What is the mass of the two isotopes of magnesium shown in Model 1? How many protons and neutrons are in each?
Each Mg-24 atom has 12 protons and 12 neutrons. Each Mg-26 atom has 12 protons and 14 neutrons.
- As a group, calculate the average mass of the two isotopes of magnesium shown in Model 1. How many protons and neutrons are in each?

Periodicity Of Elements Pogil Key

Tao Wei



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