

# Light & Periodicity HOMEWORK

## Lesson 1: Homework

1. Match the correct term, listed to EACH symbol below:

- a.  $2.998 \times 10^8$   
 b. wavelength  
 c. frequency  
 d. energy  
 e.  $2.998 \times 10^{17}$

- a.  $6.626 \times 10^{-34}$   
 b. frequency  
 c. energy  
 d.  $2.998 \times 10^8$   
 e. wavelength

2. Identify the units for EACH symbol below:

- a.  $c$  (m/s)  
 b.  $\lambda$  (m)  
 c.  $\nu$  (sec<sup>-1</sup>)

- a. meters (m)  
 b. 1/photon  
 c. sec<sup>-1</sup>  
 d. nanometers (nm)  
 e. 1/m  
 f. 1/m<sup>2</sup>

3. Convert  $6.85 \times 10^2$  nanometers to meters.

$$6.85 \times 10^2 \text{ nm} \times (1 \text{ m} / 10^9 \text{ nm}) = 6.85 \times 10^{-7} \text{ m}$$

4. As you orbit during orbit orbit around the sun, there is  $1.49 \times 10^8$  meters from Earth.

a) If a signal, which uses light waves, is sent that distance (a wavelength of  $1.49 \times 10^8$  meters), what is its frequency?

$$\begin{aligned} c &= \lambda \nu \\ (2.998 \times 10^8 \text{ m/s}) &= (1.49 \times 10^8 \text{ m}) \nu \\ \nu &= 2.01207986 \times 10^0 \text{ sec}^{-1} \end{aligned}$$

b) The units on the frequency is sec<sup>-1</sup>. If this value is inverted, the units are inverted as well. How many seconds does it take to send this signal? Give your answer with units from part "a" in this calculation.

$$(1 / 2.01207986 \text{ sec}^{-1}) = 0.4970154 \text{ sec}$$

5. Light can be used to excite the electrons in atoms. Certain frequencies of light work best for specific atoms. To excite an atom of Calcium, the ideal frequency of  $5.2 \times 10^{14} \text{ s}^{-1}$  is needed. What is the wavelength (in nm) of this electromagnetic radiation? Give your answer in scientific notation and with units.

$$\begin{aligned} c &= \lambda \nu \\ (2.998 \times 10^8 \text{ m/s}) &= (5.2 \times 10^{14} \text{ s}^{-1}) \lambda \\ \lambda &= 5.7653846 \times 10^{-7} \text{ m} \end{aligned}$$

$$5.7653846 \times 10^{-7} \text{ m} \times (10^9 \text{ nm} / 1 \text{ m}) = 576.53846 \text{ nm} \approx 576.54 \text{ nm}$$

6. The energy required to dissociate the  $\text{Cl}_2$  molecule into two separate Cl atoms requires a wavelength of  $5.00 \times 10^2$  meters. What is the energy of this light in joules?

$$\begin{aligned} (2.998 \times 10^8 \text{ m/s}) &= (5.00 \times 10^2 \text{ m}) \nu \\ \nu &= 5.996 \times 10^{14} \text{ s}^{-1} \\ E &= h \nu \\ E &= (6.626 \times 10^{-34} \text{ J} \cdot \text{s}) (5.996 \times 10^{14} \text{ s}^{-1}) \\ E &= 3.973 \times 10^{-19} \text{ J} \end{aligned}$$

7. Refrigerator coolant systems require compounds known as chlorofluorocarbons (CFCs). If not properly disposed, a refrigerator can leak this gas into the environment. Thus, ultraviolet radiation from the sun can destroy these CFCs, which can then destroy the ozone layer. How much energy do the CFCs receive from UV light that has a frequency of  $6.85 \times 10^{14} \text{ s}^{-1}$ . Give your answer in joules.

$$\begin{aligned} E &= h \nu \\ E &= (6.626 \times 10^{-34} \text{ J} \cdot \text{s}) (6.85 \times 10^{14} \text{ s}^{-1}) \\ E &= 4.54 \times 10^{-19} \text{ J} \end{aligned}$$

8. You want your lawn sprinklers to come on in the morning, at first light. So, you design a switch that will operate by the photoelectric effect when enough sunlight hits it. As sunlight hits on your switch, the metal inside the device absorbs energy and ejects electrons to create electricity that powers the sprinklers. If the switch requires  $5.8 \times 10^{14} \text{ s}^{-1}$  (just off metal to remove an electron), will the switch work if the sunlight has a wavelength of 540 nm?

$$\begin{aligned} c &= \lambda \nu \\ (2.998 \times 10^8 \text{ m/s}) &= (540 \text{ nm}) \nu \\ \nu &= 5.55185185 \times 10^{14} \text{ s}^{-1} \end{aligned}$$

$$\begin{aligned} c &= \lambda \nu \\ (2.998 \times 10^8 \text{ m/s}) &= (540 \text{ nm}) \nu \\ \nu &= 5.55185185 \times 10^{14} \text{ s}^{-1} \end{aligned}$$

$$\begin{aligned} E &= h \nu \\ E &= (6.626 \times 10^{-34} \text{ J} \cdot \text{s}) (5.55185185 \times 10^{14} \text{ s}^{-1}) \\ E &= 3.68 \times 10^{-19} \text{ J} \end{aligned}$$

$$\begin{aligned} E &= h \nu \\ E &= (6.626 \times 10^{-34} \text{ J} \cdot \text{s}) (5.55185185 \times 10^{14} \text{ s}^{-1}) \\ E &= 3.68 \times 10^{-19} \text{ J} \end{aligned}$$

# Periodicity Ap Chemistry Answers

**Clifford Lane**



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AP® Chemistry Crash Course, For the 2020 Exam, Book + Online Adrian Dingle, 2020-02-04 This Crash Course book is designed to be an aide for someone working to pass the AP advanced placement test in Chemistry

**CliffsNotes AP Chemistry** Bobrow Test Preparation Services, 2009-02-09 The book itself contains chapter length subject reviews on every subject tested on the AP Chemistry exam as well as both sample multiple choice and free response questions at each chapter's end Two full length practice tests with detailed answer explanations are included in the book

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TABLE OF CONTENT

PART I Based on Class XI NCERT

Some Basic Concepts of Chemistry Structure of Atom Classification of Elements and Periodicity in Properties Chemical Bonding and Molecular Structure States of Matter Thermodynamics Equilibrium Redox Reactions Hydrogen s Block Elements p Block Elements Organic Chemistry Some Basic Principles and Techniques Hydrocarbons Environmental Chemistry

PART II Based on Class XII NCERT

The Solid State Solutions Electrochemistry Chemical Kinetics Surface Chemistry Nuclear Chemistry p Block Elements The d and f Block Elements Coordination Compounds Haloalkanes and Haloarenes Alcohols Phenols and Ethers Aldehydes Ketones and Carboxylic Acids Nitrogen Containing Compounds Biomolecules Polymers Chemistry in Everyday Life Analytical Chemistry General Principles and Processes of Isolation of Elements

Questions Asked in JEE Main 2015 Solved Papers 2016 JEE Main BITSAT AP EAMCET TS EAMCET GGSIPU Solved Papers 2017 JEE Main Advanced BITSAT VIT WBJEE Solved Papers 2018 JEE Main Advanced BITSAT WBJEE Solved Papers 2019 JEE Main Advanced BITSAT WBJEE

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## **Table of Contents Periodicity Ap Chemistry Answers**

1. Understanding the eBook Periodicity Ap Chemistry Answers
  - The Rise of Digital Reading Periodicity Ap Chemistry Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Periodicity Ap Chemistry Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Periodicity Ap Chemistry Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Periodicity Ap Chemistry Answers
  - Personalized Recommendations
  - Periodicity Ap Chemistry Answers User Reviews and Ratings

- Periodicity Ap Chemistry Answers and Bestseller Lists
- 5. Accessing Periodicity Ap Chemistry Answers Free and Paid eBooks
  - Periodicity Ap Chemistry Answers Public Domain eBooks
  - Periodicity Ap Chemistry Answers eBook Subscription Services
  - Periodicity Ap Chemistry Answers Budget-Friendly Options
- 6. Navigating Periodicity Ap Chemistry Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Periodicity Ap Chemistry Answers Compatibility with Devices
  - Periodicity Ap Chemistry Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Periodicity Ap Chemistry Answers
  - Highlighting and Note-Taking Periodicity Ap Chemistry Answers
  - Interactive Elements Periodicity Ap Chemistry Answers
- 8. Staying Engaged with Periodicity Ap Chemistry Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Periodicity Ap Chemistry Answers
- 9. Balancing eBooks and Physical Books Periodicity Ap Chemistry Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Periodicity Ap Chemistry Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Periodicity Ap Chemistry Answers
  - Setting Reading Goals Periodicity Ap Chemistry Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Periodicity Ap Chemistry Answers
  - Fact-Checking eBook Content of Periodicity Ap Chemistry Answers
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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