



Percent Composition Of Compounds

DATE _____

Find the percent composition in the following compounds.

1. H_2O

$$\text{molar mass: } 2 \times 1.01 + 16.00 = 18.02 \text{ g mol}^{-1}$$

$$\% \text{H} = \frac{2 \times 1.01}{18.02} \times 100\% = 11.2\%$$

$$\% \text{O} = \frac{16.00}{18.02} \times 100\% = 88.8\%$$

2. CO_2

$$\text{molar mass: } 12.01 + 2 \times 16.00 = 44.01 \text{ g mol}^{-1}$$

$$\% \text{C} = \frac{12.01}{44.01} \times 100\% = 27.3\%$$

$$\% \text{O} = \frac{2 \times 16.00}{44.01} \times 100\% = 72.7\%$$

3. NaOH

$$\text{molar mass: } 22.99 + 16.00 + 1.01 = 40.00 \text{ g mol}^{-1}$$

$$\% \text{Na} = \frac{22.99}{40.00} \times 100\% = 57.5\%$$

$$\% \text{H} = \frac{1.01}{40.00} \times 100\% = 2.5\%$$

$$\% \text{O} = \frac{16.00}{40.00} \times 100\% = 40.0\%$$

4. H_2SO_4

$$\text{molar mass: } 2 \times 1.01 + 32.07 + 4 \times 16.00 = 98.09 \text{ g mol}^{-1}$$

$$\% \text{H} = \frac{2 \times 1.01}{98.09} \times 100\% = 2.1\%$$

$$\% \text{S} = \frac{32.07}{98.09} \times 100\% = 32.7\%$$

$$\% \text{O} = \frac{4 \times 16.00}{98.09} \times 100\% = 65.2\%$$

5. Ca(OH)_2

$$\text{molar mass: } 40.08 + 2 \times 16.00 + 2 \times 1.01 = 74.10 \text{ g mol}^{-1}$$

$$\% \text{Ca} = \frac{40.08}{74.10} \times 100\% = 54.1\%$$

$$\% \text{H} = \frac{2 \times 1.01}{74.10} \times 100\% = 2.7\%$$

$$\% \text{O} = \frac{2 \times 16.00}{74.10} \times 100\% = 43.2\%$$

Percentage Composition Chemistry Answerkey

Lei Huang

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Table of Contents Percentage Composition Chemistry Answerkey

1. Understanding the eBook Percentage Composition Chemistry Answerkey
 - The Rise of Digital Reading Percentage Composition Chemistry Answerkey
 - Advantages of eBooks Over Traditional Books
2. Identifying Percentage Composition Chemistry Answerkey
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Percentage Composition Chemistry Answerkey
 - User-Friendly Interface
4. Exploring eBook Recommendations from Percentage Composition Chemistry Answerkey
 - Personalized Recommendations
 - Percentage Composition Chemistry Answerkey User Reviews and Ratings
 - Percentage Composition Chemistry Answerkey and Bestseller Lists
5. Accessing Percentage Composition Chemistry Answerkey Free and Paid eBooks
 - Percentage Composition Chemistry Answerkey Public Domain eBooks
 - Percentage Composition Chemistry Answerkey eBook Subscription Services
 - Percentage Composition Chemistry Answerkey Budget-Friendly Options

6. Navigating Percentage Composition Chemistry Answerkey eBook Formats
 - ePub, PDF, MOBI, and More
 - Percentage Composition Chemistry Answerkey Compatibility with Devices
 - Percentage Composition Chemistry Answerkey Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Percentage Composition Chemistry Answerkey
 - Highlighting and Note-Taking Percentage Composition Chemistry Answerkey
 - Interactive Elements Percentage Composition Chemistry Answerkey
8. Staying Engaged with Percentage Composition Chemistry Answerkey
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Percentage Composition Chemistry Answerkey
9. Balancing eBooks and Physical Books Percentage Composition Chemistry Answerkey
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Percentage Composition Chemistry Answerkey
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Percentage Composition Chemistry Answerkey
 - Setting Reading Goals Percentage Composition Chemistry Answerkey
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Percentage Composition Chemistry Answerkey
 - Fact-Checking eBook Content of Percentage Composition Chemistry Answerkey
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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