

## 11-1 Practice Problems

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1. Lead will react with Hydrochloric acid to produce Lead (II) Chloride and Hydrogen gas. How many moles of Hydrochloric acid are needed to completely react with 0.36-mol of lead?

2. How many moles of  $\text{HNO}_3$  will be produced when 0.51-mol of  $\text{N}_2\text{O}_5$  reacts according to the following equation?  
 $\text{N}_2\text{O}_5 + \text{H}_2\text{O} \rightarrow 2\text{HNO}_3$

3. Iron will react with Oxygen Gas to produce  $\text{Fe}_2\text{O}_3$ . How many moles of  $\text{Fe}_2\text{O}_3$  will be produced if 0.18-mol of Fe reacts?

4. How many moles of NaBr will be produced when 0.69-mol of Bromine reacts according to the following equation?  
 $\text{Br}_2 + 2\text{NaI} \rightarrow 2\text{NaBr} + \text{I}_2$

5. Phosphorus will react with Bromine to produce Phosphorus Tribromide. How many moles of Phosphorus Tribromide will be produced if 0.78-mol of Bromine is reacted?

6. Carbon will react with Zinc Oxide to produce Zinc and Carbon Dioxide. How many moles of carbon dioxide will be produced if 0.38-mol of  $\text{ZnO}$  is completely reacted?

7. How many moles of oxygen will be needed to react with 0.38-mol of  $\text{C}_3\text{H}_8$  according to the following equation?  
 $\text{C}_3\text{H}_8 + 5\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$

8. Nitrogen gas can react with Hydrogen gas to produce ammonia. How many moles of Nitrogen will be needed to produce 0.48-mol of  $\text{NH}_3$ ?

9. How many moles of hydrogen will be produced if 0.44-mol of  $\text{CaH}_2$  reacts according to the following equation?  
 $\text{CaH}_2 + 2\text{H}_2\text{O} \rightarrow \text{Ca(OH)}_2 + 2\text{H}_2$

10. How many moles of water will be produced if 2.35-mol of oxygen reacts according to the following equation?  
 $2\text{C}_6\text{H}_6 + 15\text{O}_2 \rightarrow 12\text{CO}_2 + 6\text{H}_2\text{O}$

# Prentice Hall Chemistry Chapter 1 practice Problems

**Robert R. Redfield**



**Prentice Hall Chemistry Chapter 1practice Problems:**

## Decoding **Prentice Hall Chemistry Chapter 1practice Problems**: Revealing the Captivating Potential of Verbal Expression

In a time characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its power to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Prentice Hall Chemistry Chapter 1practice Problems**," a mesmerizing literary creation penned with a celebrated wordsmith, readers set about an enlightening odyssey, unraveling the intricate significance of language and its enduring impact on our lives. In this appraisal, we shall explore the book's central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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