

14

THE BEHAVIOR OF GASES

Chapter Test II

A. Matching

Match a word or phrase in Column I with the correct description in Column II. Write the letter of the correct answer next to the item.

Column I

1. The constant relating unit temperature changes to changes in volume for a fixed amount of gas at constant pressure (the ratio of the change in volume to the change in temperature).
2. The volume of a fixed amount of a gas at constant pressure and temperature when the pressure is 1 atm.
3. The ratio of the mass of a gas to its volume, independent of the actual amount of gas in the sample.
4. An experimental gas constant that varies with the pressure of the sample.
5. A constant of nature that has values of about 8.314 J/mol·K.
6. The a parameter of gas constant equations for monatomic gas molecules only (about 1.67).
7. An experimental gas constant that is about 0.0821 L·atm/mol·K.
8. $\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$
9. The amount of heat enough to raise 1 g of water 1°C.
10. A unit of energy (Joules).

Column II

- a. pressure-volume work
- b. ideal gas constant (R)
- c. volume
- d. compressibility
- e. Boyle's law
- f. molar volume
- g. standard molar volume (at STP)
- h. density
- i. Joules/mole
- j. heat capacity of oxygen

B. Multiple Choice

Choose the correct answer and write its letter on the line.

1. Subliming the solvent of a solid sample for mass determination (drying) requires constant pressure and temperature.
a. the atmosphere and 0°C
b. the vacuum of a desiccator
c. the atmosphere and room temp.
d. the vacuum of a desiccator

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