

11-2 Practice Problems

1. Determine the mass of lithium hydroxide produced when 0.38 g of lithium nitride reacts with water according to the following equation:



$$\frac{0.38\text{g Li}_3\text{N}}{1} \times \frac{1\text{mol Li}_3\text{N}}{35\text{g Li}_3\text{N}} \times \frac{3\text{mol LiOH}}{1\text{mol Li}_3\text{N}} \times \frac{24\text{g LiOH}}{1\text{mol LiOH}} = \boxed{0.78\text{g LiOH}}$$

2. What mass of sodium chloride is produced when chlorine reacts with 0.29 g of sodium iodide?



$$\frac{0.29\text{g NaI}}{1} \times \frac{1\text{mol}}{150} \times \frac{2\text{mol NaCl}}{2\text{mol NaI}} \times \frac{58\text{g}}{1\text{mol}} = \boxed{0.112\text{g NaCl}}$$

3. Determine the mass of carbon dioxide produced when 0.85 g of butane reacts with oxygen according to the following equation:



$$\frac{0.85\text{g C}_4\text{H}_{10}}{1} \times \frac{1\text{mol}}{58\text{g}} \times \frac{8\text{mol CO}_2}{2\text{mol C}_4\text{H}_{10}} \times \frac{44\text{g}}{1\text{mol}} = \boxed{2.58\text{g CO}_2}$$

4. Determine the mass of antimony produced when 0.46 g of antimony(III) oxide reacts with carbon according to the following equation:



$$\frac{0.46\text{g Sb}_2\text{O}_3}{1} \times \frac{1\text{mol}}{291.52\text{g}} \times \frac{2\text{mol Sb}}{1\text{mol Sb}_2\text{O}_3} \times \frac{121.76\text{g}}{1\text{mol}} = \boxed{0.38\text{g Sb}}$$

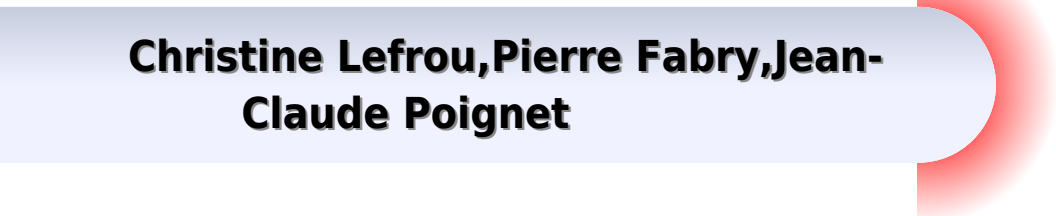
5. What mass of hydrogen peroxide (H_2O_2) must decompose to produce 0.77 g of water?



$$\frac{0.77\text{g H}_2\text{O}}{1$$

Prentice Hall Chemistry Chapter 11 Answer Key

**Christine Lefrou, Pierre Fabry, Jean-
Claude Poignet**



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