

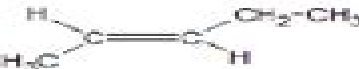
A-LEVEL CHEMISTRY

MARK SCHEME

Practice Paper 2

Maximum marks: 101

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Q1.	(a)	(i)	M1 (Compounds / molecules with) the <u>same structural formula</u> <i>Penalise M1 if 'same structure' or 'different structural / displayed formula'.</i> M2 with atoms / bonds / groups arranged differently in space OR atoms / bonds / groups with different spatial arrangements / different orientation <i>Ignore references to 'same molecular formula' or 'same empirical formula'.</i> <i>Mark independently.</i>	2
		(ii)	 <i>Credit C-H₃C</i> <i>Credit C₂H₄</i> <i>Penalise C-CH₃CH₃</i>	1
	(b)		M1 Br₂ OR bromine (water) OR bromine (in CCl₄ / organic solvent) <i>If M1, has no reagent or an incorrect reagent, CE=0.</i> <i>Ignore 'acidified'.</i> M2 Isomer 1: decolourised / goes colourless / loses its colour <i>For M1 penalise Br (or incorrect formula of other correct reagent), but mark on.</i> M3 Isomer 2: remains orange / red / yellow / brown / the same OR no reaction / no (observable) change OR reference to colour going to the cyclopentane layer <i>For M1, it must be a whole reagent and / or correct formula.</i> <i>If oxidation state given in name, it must be correct. If 'manganate' OR 'manganate(IV)' or incorrect formula, penalise M1, but mark on.</i>	1 1 1
			Alternatives : potassium manganate(VII)	
			M1 KMnO₄ in acid M2 colourless M3 purple	
			M1 KMnO₄ in alkali / neutral M2 brown solid M3 purple	
			Credit for the use of iodine	
			M1 iodine (solution / in KI) M2 colourless M3 (brown) to purple (credit no change)	
			Credit for the use of <u>concentrated</u> H₂SO₄	

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