

Your instructor may ask you to repeat one of your trials if your error is greater than 5%.

INTERPRETATION OF DATA

Your data will be used to calculate the theoretical volume of the air in the experiment during the various temperature changes using Charles' Law. The volume of hot air is measured directly by filling the flask with water up to the wax pencil mark and measuring the volume of water (#4, data sheet). The volume of the air after cooling is found by subtracting the volume of water remaining in the cooled flask (#3, data sheet) from the volume held by the full flask (#4). The volume of water sucked into the flask upon cooling represents the extent of contraction of the hot air upon cooling.

Charles' Law can be expressed as:

$$V_{\text{hot gas}} = \text{constant} \times T_{\text{hot gas}} ; \quad V_{\text{cold gas}} = \text{constant} \times T_{\text{cold gas}}$$

or

$$\frac{V_{\text{cold}}}{V_{\text{hot}}} = \frac{T_{\text{cold}}}{T_{\text{hot}}}$$

The temperatures used must be on the Kelvin temperature scale ($K = ^\circ\text{C} + 273.15$). Since your temperature readings have a precision of $\pm 0.1^\circ\text{C}$, add 273.15 but then round to one decimal place. Blanks have been provided on the data sheet for entering these values. From the formula shown above,

$$V_{\text{cold}} = \left(\frac{T_{\text{cold}}}{T_{\text{hot}}} \right) V_{\text{hot}}$$

Use this expression to calculate what the volume of cold air should be.

Calculate the percentage difference between the calculated and measured values as follows:

$$\% \text{ difference} = \frac{(\text{calculated volume of cold air} - \text{measured volume of cold air})}{(\text{calculated volume of cold air})} \times 100 \%$$

Your prediction for Absolute Zero can be found by plotting the volume of gas vs. temperature and extrapolating to zero volume (remember that the molecules of an ideal gas occupy no space). The temperature at zero volume is absolute zero.

To do this, prepare a computer-generated scatter plot of the data with temperature on the y-axis and volume on the x-axis. There will be four data points. Do not connect the dots.

Data point	Volume, mL	Temperature, $^\circ\text{C}$
1	hot air, all trials averaged (#5)	boiling water, all trials averaged (#1)
2	cold air, trial 1 (#6)	ice bath, trial 1 (#2)
3	cold air, trial 2 (#6)	tap water, trial 2 (#2)
4	cold air, trial 3 (#6)	warm water, trial 3 (#2)

Have the program add a best fit line through the data and display the equation of this line. The y-intercept of this equation is your prediction of Absolute Zero from your data. An example follows.

Prentice Hall Chemistry 110 Lab Manual Answer Key

Yicheng Fang



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