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Name \_\_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_

17

#### THERMOCHEMISTRY

#### SECTION 17.1 THE FLOW OF ENERGY—HEAT AND WORK (pages 505-510)

This section explains the relationship between energy and heat, and distinguishes between heat capacity and specific heat.

#### ▶ Energy Transformations (page 505)

- L. What area of study in chemistry is concerned with the heat transfers that occur during chemical reactions? <a href="https://doi.org/10.1007/j.com/instry-10.1007/j.com/inst
- Where the use of energy is concerned (in a scientific sense), when is work done?Work is done when a force is used to move an object.
- 3. Circle the letter next to each sentence that is true about energy.
- (a.) Energy is the capacity for doing work or supplying heat.
- (b.) Energy is detected only because of its effects.
- c. Heat is energy that transfers from one object to another because they are at the same temperature.
- (d) Gasoline contains a significant amount of chemical potential energy.
- 4. Circle the letter next to each sentence that is true about heat.
  - a. One effect of adding heat to a substance is an increase in the temperature of that substance.
  - Heat always flows from a cooler object to a warmer object.
  - (c) If two objects remain in contact, heat will flow from the warmer object to the cooler object until the temperature of both objects is the same.

#### Exothermic and Endothermic Processes (pages 506–507)

What can be considered the "system" and what are the "surroundings" when studying a mixture of chemicals undergoing a reaction? Write your answers where indicated below.

System: The mixture of chemicals itself is considered the system.

Surroundings: Everything but the mixture of chemicals is the surroundings, but

practically speaking, the immediate vicinity of the system.

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