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11-3 Practice Problems

 Literary the limiting reactant when 1.22-g of O₂ reacts with 1.55-g of H₂ to produce water. Identify the limiting reactors when 65.14- g of CaCl, reacts with 74.65-g of Na-CO₃ to produce CaCO₃ and NaCl.

 Mentify the limiting reaction when 4.68-g of Percects with 2.88-g of S to produce FeS. Identify the limiting reactant when 4.687-g of SF₄ reacts with 6.281-g of I₂O₃ to produce IF₂ and SO₃.

 Identify the limiting reactant when 5.87-g of Mg(OH)₃ means with 12.84-g of HC1 to form MgCl₂ and water. 10. If 4.1-g of Cr is heated with 9.3-g of Cl₂, what mass CrCl₃ will be produced?

 Identify the limiting reactant when 6.25-g of AgNO₃ reacts with 4.12-g of NaCl to form NaNO₃ and AgCl. What mass of SO₁ is produced from the reaction between 31.5-g of S₀ and 8.65-g of O₂?

 Identify the limiting resented when T.B.L-g of HICL reacts with 3.24-g of NaOH to produce NaCL and HigO. 12. What wass of SO₃ is produced from the reaction of 12.4-g of SO₃ and 3.45-g of O₃?

 Identify the limiting reactant when 6.33-g of H₂SO₄ reacts with 5.92-g of NaOH to produce Na₂SO₄ and water. What mass of SI₂SO₂ is produced from the reaction of 6.58-g of SO₂ and 1.64-g of HsO?

 Identify the limiting reactant when 43.25-g of CaC₂ reacts with 33.71-g of water to produce CatOH₁, and CaH₂. 14. What mass of CdS is produced if 8.47-g of cadmium reacts with 2.51-g of Saltur?

Prentice Hall Chemistry Practice Problem

Michael Brown

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