

What are Acids, Bases, and Salts?

The Swedish chemist Svante Arrhenius introduced the theory of ionization and used this theory to explain much about the behavior of acids and bases.

- An Arrhenius acid is defined as any compound that dissociates in aqueous solution to form H^+ ions.



- An Arrhenius base is defined as any compound that dissociates in aqueous solution to form OH^- ions.



- Salts are compounds that dissociate in aqueous solution releasing neither H^+ ions nor OH^- ions.



Using the **Arrhenius definition**, classify the following examples as acids, bases, or salts:

HBr	<u>acid</u>	KCl	<u>salt</u>
Mg(OH) ₂	<u>base</u>	H ₃ PO ₄	<u>acid</u>
HCl	<u>acid</u>	HClO	<u>acid</u>
KNO ₃	<u>salt</u>	Al(OH) ₃	<u>base</u>
HFO ₄	<u>acid</u>	KC ₂ H ₃ O ₂	<u>salt</u>
Ba(OH) ₂	<u>base</u>	NaCl	<u>salt</u>

Note Taking Worksheet Acids Bases And Salts

Hussin A.Rothana



Note Taking Worksheet Acids Bases And Salts:

Unveiling the Power of Verbal Artistry: An Mental Sojourn through **Note Taking Worksheet Acids Bases And Salts**

In a world inundated with monitors and the cacophony of immediate communication, the profound energy and emotional resonance of verbal artistry usually diminish in to obscurity, eclipsed by the constant assault of noise and distractions. Yet, nestled within the lyrical pages of **Note Taking Worksheet Acids Bases And Salts**, a fascinating function of literary splendor that pulses with organic emotions, lies an memorable journey waiting to be embarked upon. Written by a virtuoso wordsmith, this magical opus manuals viewers on a mental odyssey, softly exposing the latent possible and profound affect stuck within the elaborate web of language. Within the heart-wrenching expanse of the evocative examination, we can embark upon an introspective exploration of the book is central styles, dissect their charming publishing model, and immerse ourselves in the indelible impact it leaves upon the depths of readers souls.

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