

Question number	Answer	Marks	Guidance															
1 (a)	C	B1																
1 (b)	B	B1																
1 (c)	D	B1																
2 (a)	38	B1																
2 (b)	32	B1																
2 (c)	5	B1																
3 (a) (i)	23	B1																
3 (a) (ii)	$^{54}_{26}\text{Fe}$	B2																
3 (b)	$^{24}_{11}\text{Na}$	B2																
3 (c)	$^{88}_{38}\text{Sr}$	B2																
4 (a)	Isotopes are atoms of the same element with different numbers of neutrons and different masses.	B1																
4 (b)	Relative isotopic mass is the mass of an isotope relative to one-twelfth of the mass of an atom of carbon-12.	B2																
4 (c)	71	B1																
4 (d)	<table><tr><td></td><td>protons</td><td>neutrons</td><td>electrons</td><td>charge</td></tr><tr><td>^{17}O</td><td>8</td><td>9</td><td>10</td><td>2-</td></tr><tr><td>^{27}Al</td><td>13</td><td>14</td><td>10</td><td>3+</td></tr></table>		protons	neutrons	electrons	charge	^{17}O	8	9	10	2-	^{27}Al	13	14	10	3+	B2	
	protons	neutrons	electrons	charge														
^{17}O	8	9	10	2-														
^{27}Al	13	14	10	3+														
5 (a)	Relative atomic mass is the weighted mean mass of an atom of an element relative to one-twelfth of the mass of an atom of carbon-12.	B3																
5 (b)	$A_r = 32.09$	B2																
5 (c)	<table><tr><td></td><td>protons</td><td>neutrons</td><td>electrons</td></tr><tr><td>^{32}S</td><td>16</td><td>16</td><td>16</td></tr><tr><td>$^{32}\text{S}^{2-}$</td><td>16</td><td>16</td><td>18</td></tr></table>		protons	neutrons	electrons	^{32}S	16	16	16	$^{32}\text{S}^{2-}$	16	16	18	B1				
	protons	neutrons	electrons															
^{32}S	16	16	16															
$^{32}\text{S}^{2-}$	16	16	18															
6 (a)	$3\text{Mg(s)} + \text{N}_2\text{(g)} \rightarrow \text{Mg}_3\text{N}_2\text{(s)}$ 1 mark for correct species and state symbols 1 mark for balance	B1 × 2																
6 (b)	$\text{Ca(s)} + 2\text{H}_2\text{O(l)} \rightarrow \text{Ca(OH)}_2\text{(aq)} + \text{H}_2\text{(g)}$ 1 mark for correct species and state symbols 1 mark for balance	B1 × 2																
6 (c)	$6\text{NaOH(aq)} + \text{Fe}_2(\text{SO}_4)_3\text{(aq)} \rightarrow 2\text{Fe(OH)}_3\text{(s)} + 3\text{Na}_2\text{SO}_4\text{(aq)}$ 1 mark for correct species and state symbols 1 mark for balance	B1 × 2																

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OCR (A) AS Chemistry Mike Smith, 2008-09-16 Student Unit Guides are perfect for revision Each guide is written by an examiner and explains the unit requirements summarises the relevant unit content and includes a series of specimen questions and answers There are three sections to each guide Introduction includes advice on how to use the guide an explanation of the skills being tested by the Assessment Objectives an outline of the unit or module and depending on the unit suggestions for how to revise effectively and prepare for the examination questions Content Guidance an outline of the areas and the key terms and concepts students will need to focus on It is designed to help them structure their revision and make them aware of the concepts they need to understand the exam Question and Answers sample questions and answers for each unit exam Examples of students work to demonstrate good practice are featured together with examiner s comments explaining how the marks are achieved

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questions Offer support beyond the Student books with coverage of methodologies and generic practical skills not focused on in the textbooks

A2 Level Chemistry for OCR A: Student Book CGP Books, Richard Parsons, 2012-06-01 This student friendly textbook comprehensively covers the OCR A exam specification for A2 Level Chemistry It contains in depth accessible notes explaining every topic supported by clear diagrams photographs tips and worked examples To test students knowledge and understanding there are practice questions throughout the book with complete answers included There s also detailed guidance for success in the final assessments including exam style questions and sections dedicated to exam skills and how science works

GCSE Chemistry, 2016 **OCR AS Physics** Roger Hackett, 2007 **Revise A2 for Salters New Edition** Dave Newton, 2009-06-01 Suitable for the 2008 OCR A Level Chemistry B Salters specification this guide provides students with help they need for exam preparation It also provides comprehensive coverage of the specification for the exams bite sized chunks of information and quick check revision questions to test knowledge and understanding

New GCSE Chemistry: OCR 21st Century Answers (for Exam Practice Workbook) CGP Books, 2016 **OCR A Level Chemistry Student Book 2** Mike Smith, John Older, 2015-09-11 Exam Board OCR Level A level Subject Chemistry First Teaching September 2015 First Exam June 2017 This is an OCR endorsed resource Stretch and challenge your students knowledge and understanding of Chemistry build their mathematical and practical skills and provide plenty of assessment guidance with this OCR Year 2 Student Book Build understanding with a summary of prior knowledge and diagnostic questions at the start of each chapter to help bring students up to speed Support practical assessment with Practical Skill summaries that help develop your students knowledge and skills Test understanding and provide plenty of practice to assess progression with Test Yourself Questions and multiple choice questions Provide mathematical support with examples of method integrated throughout and a dedicated Maths in Chemistry chapter Develop understanding with free online access to Test yourself Answers an Extended Glossary Learning Outcomes and Topic Summaries

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