

GROUP 2 ELEMENTS AND THEIR COMPOUNDS

- show their outermost electron placed on a subshell
- readily react to lose 2 electrons
 - $\text{M} \rightarrow \text{M}^{2+} + 2\text{e}^{-}$

Trends down the group:

- 1st ionisation energy decreases
 - Distance between outermost electron and nucleus increases
 - More shielding
 - Counteractive effect of increasing nuclear charge
- Atomic radius increases
- Melting point generally decreases
 - Much lower than all transition metals

Physical properties:

- Good conductors of heat and electricity
 - Sea of delocalised electrons
- Form white compounds
 - Transition metals form coloured compounds
- Oxidation number of +2
- All compounds are toxic or almost toxic
 - Be compounds show some covalent character because small highly charged cation polarises anion shell of anion
- Oxides / hydroxides are basic
 - React with acids (proton acceptors)
 - Example:
 - BaO
 - $\text{BaO} + 2\text{H}^{+} \rightarrow \text{Ba}^{2+} + \text{H}_2\text{O}$, which is exothermic
- Density = 2.83 g/cm^3
 - Average density of transition metals is about 7 g/cm^3
- Compared to group 2 metals:
 - Harder + denser
 - Higher melting point
 - Metallic bond is stronger (higher charge on cation & more electrons delocalised)

Chemical properties:

- Strong reducing agents
- Oxidation is losing electrons
- Reduction is gaining electrons:
 - Reducing agent (donates electrons (reducing) other substance)
 - Oxidising agent accepts electrons (oxidising itself)
- E.g.
 - $\text{ZnO} + \text{C} \rightarrow \text{Zn} + \text{CO}_2$
 - $\text{ZnO} \rightarrow \text{Zn}^{2+} + 2\text{O}^{2-}$
 - $\text{C} + 2\text{O}^{2-} \rightarrow \text{CO}_2^{2-}$

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RC Schank



Ocr Chemistry June 2014 F321:

Embracing the Melody of Phrase: An Emotional Symphony within **Ocr Chemistry June 2014 F321**

In some sort of taken by displays and the ceaseless chatter of fast interaction, the melodic beauty and mental symphony produced by the written word often fade into the background, eclipsed by the persistent noise and distractions that permeate our lives. Nevertheless, nestled within the pages of **Ocr Chemistry June 2014 F321** a wonderful fictional treasure filled with organic emotions, lies an immersive symphony waiting to be embraced. Constructed by an outstanding composer of language, this charming masterpiece conducts viewers on a mental trip, well unraveling the hidden tunes and profound impact resonating within each cautiously crafted phrase. Within the depths with this touching evaluation, we will explore the book is central harmonies, analyze its enthralling publishing design, and surrender ourselves to the profound resonance that echoes in the depths of readers souls.

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