

Page 2	Mark Scheme: Teachers' version IGCSE – May/June 2011	Syllabus 0620	Paper 31
	(c) base not alkali [1]		
	accepts a proton accepts hydrogen ion / H^+ only [1]		[2]
	proton and H^+ [2]		
3	(a) any four max 4 carbon forms carbon dioxide / carbon monoxide [1] this is a gas it escapes / blown out / diffuses [1] silicon forms silicon(IV) oxide / silica [1] / silicon(IV) oxide present in impure iron silicon(IV) oxide reacts with calcium oxide to form slag or calcium silicate [1] slag removed from surface [1] accept skimmed, syphoned, poured off not tapped max [4] accept correct formula or equations not calcium oxide reacts with silicon		
	(b) (i) any sensible suggestion – harder/stronger/can be tailored for a specific use/more resistant to corrosion [1] not steel does not rust		
	(ii) mild steel – cars or any vehicle/bicycles/white goods/screws or nails/poll/bridges/boots/ buildings/ships/pipes/machinery etc. [1] stainless steel – chemical plants/cooking utensils/jewellery/cutlery/surgical equipment/ kitchen sinks/pipes/etc. [1]		
	(c) (i) strong attractive forces / strong bonds / bonds hard to break / requires a lot of energy to break bonds [1] not between ions, not between positive and negative ions, not between electrons between positive ions and (negative) electrons / opposite charges attract [1]		
	(ii) because the layers, lattice or rows of ions/cations accept sheets of ions [1] not atoms / molecules / protons / nuclei can move / slip / slide past each other [1]		
4	(a) (i) $2ZnS + 3O_2 \rightarrow 2ZnO + 2SO_2$ [2] not balanced only [1]		
	(ii) two reagents from named metal(s) more reactive than zinc/carbon monoxide not hydrogen [2]		
	(iii) they have different boiling points [1] cadmium will distil first then zinc leaving lead/lead distilled last [1]		

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