

Boyle's Law

- The volume of a fixed mass of gas varies inversely with the pressure of constant temperature.
- math
- math

Kinetic Theory and Boyle's Law

- Pressure of a gas is caused by the collisions of the gas molecules impinging on the walls.
- If the gas is compressible to small the volume is small smaller smaller the pressure is large larger larger.
- Force exerting pressure on walls of the vessel.
- Weight of the gas will increase.

Ex. 1. A certain quantity of gas has a volume of 800 ml. at standard atmospheric pressure. After the volume is reduced, it maintains a pressure of 1.25 atm above the pressure inside the tube. What is the volume of the gas at this pressure?

$$P_1 V_1 = P_2 V_2 \quad 1 \text{ atm} \times 800 \text{ ml} = 1.25 \text{ atm} \times V_2 \quad V_2 = \frac{800 \text{ ml}}{1.25} = 640 \text{ ml}$$

Ex. 2. A gas at pressure of 1.25 atm and 400 ml. of volume is compressed to 300 ml. If the pressure is increased until the volume is 200 ml. what is the new pressure, assuming the temperature remains constant?

$$1.25 \text{ atm} \times 400 \text{ ml} = P_2 \times 200 \text{ ml} \quad P_2 = \frac{1.25 \times 400}{200} = 2.5 \text{ atm}$$

$$P_2 = 2.5 \text{ atm}$$

Note Taking Guide Episode 902 Chemistry

JR Anderson



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