

Chapter 6 Chemical Bonds

Section 6.2 Covalent Bonding**(pages 165–169)**

This section discusses the formation of covalent bonds and the factors that determine whether a molecule is polar or nonpolar. It also discusses attractions between molecules.

Reading Strategy (page 165)

Relating Text and Visuals As you read the section, look closely at Figure 9. Complete the table by describing each type of model shown. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.

Molecular Models	
Model	Description
Electron dot	
Structural formula	
Space-filling	
Electron cloud	

Covalent Bonds (pages 165–167)

- Describe a covalent bond. _____

- Circle the letters of molecular models that show orbitals of atoms overlapping when a covalent bond forms.
 - electron dot
 - structural formula
 - space-filling
 - electron cloud
- Describe a molecule. _____

- Is the following sentence true or false? In a covalent bond, the atoms are held together by the attractions between the shared electrons and the protons in each nucleus. _____
- Circle the correct answer. Nitrogen has five valence electrons. How many pairs of electrons must two nitrogen atoms share in order for each atom to have eight valence electrons?
 - zero
 - one
 - two
 - three

Note Taking Guide Covalent Bonding Answer Key

Ian Pickup



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Table of Contents Note Taking Guide Covalent Bonding Answer Key

1. Understanding the eBook Note Taking Guide Covalent Bonding Answer Key
 - The Rise of Digital Reading Note Taking Guide Covalent Bonding Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Note Taking Guide Covalent Bonding Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Note Taking Guide Covalent Bonding Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Note Taking Guide Covalent Bonding Answer Key
 - Personalized Recommendations
 - Note Taking Guide Covalent Bonding Answer Key User Reviews and Ratings
 - Note Taking Guide Covalent Bonding Answer Key and Bestseller Lists
5. Accessing Note Taking Guide Covalent Bonding Answer Key Free and Paid eBooks

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 - Note Taking Guide Covalent Bonding Answer Key eBook Subscription Services
 - Note Taking Guide Covalent Bonding Answer Key Budget-Friendly Options
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- Adjustable Fonts and Text Sizes of Note Taking Guide Covalent Bonding Answer Key
 - Highlighting and Note-Taking Note Taking Guide Covalent Bonding Answer Key
 - Interactive Elements Note Taking Guide Covalent Bonding Answer Key
8. Staying Engaged with Note Taking Guide Covalent Bonding Answer Key
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Note Taking Guide Covalent Bonding Answer Key
9. Balancing eBooks and Physical Books Note Taking Guide Covalent Bonding Answer Key
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Note Taking Guide Covalent Bonding Answer Key
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Note Taking Guide Covalent Bonding Answer Key
- Setting Reading Goals Note Taking Guide Covalent Bonding Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Note Taking Guide Covalent Bonding Answer Key
- Fact-Checking eBook Content of Note Taking Guide Covalent Bonding Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

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14. Embracing eBook Trends

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