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Chemical Equations and Stoichiometry

STOICHIOMETRY PROBLEMS

General Stoichiometry

1. Several brands of antacid tablets use aluminum hydroxide to neutralize excess acid.



(a) What quantity of HCl, in grams, can a tablet with 0.750 g of $\text{Al}(\text{OH})_3$ consume?

Handwritten calculation for part (a):

$$\frac{0.750 \text{ g Al}(\text{OH})_3}{78.01 \text{ g Al}(\text{OH})_3} \times \frac{1 \text{ mol Al}(\text{OH})_3}{1 \text{ mol Al}(\text{OH})_3} \times \frac{3 \text{ mol HCl}}{1 \text{ mol Al}(\text{OH})_3} \times \frac{36.46 \text{ g HCl}}{1 \text{ mol HCl}} = 1.03 \text{ g HCl}$$

(b) What quantity, in grams, of water is produced?

Handwritten calculation for part (b):

$$\frac{0.750 \text{ g Al}(\text{OH})_3}{78.01 \text{ g Al}(\text{OH})_3} \times \frac{1 \text{ mol Al}(\text{OH})_3}{1 \text{ mol Al}(\text{OH})_3} \times \frac{3 \text{ mol H}_2\text{O}}{1 \text{ mol Al}(\text{OH})_3} \times \frac{18.02 \text{ g H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = 0.82 \text{ g H}_2\text{O}$$

On Your Own Basic Stoichiometry

R Sanford



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