

- 1.1 Find the rate of reaction defined as

$$r_{O_2} = \frac{\text{mol } O_2 \text{ used}}{\text{sec} \cdot \text{m}^3 \text{ of tank}}$$

Evaluate terms

$$\bar{t} = \frac{V}{\bar{v}} \quad \text{or} \quad V = \bar{t} \bar{v}$$

or (Volume of treatment tanks) = $\left(\frac{1}{3} \text{ day}\right) \left(32000 \frac{\text{m}^3}{\text{day}}\right) = 10667 \text{ m}^3$

O_2 used
 $\left(200 \frac{\text{mg}}{\text{lit}}\right) \left(\frac{1 \text{ gm}}{1000 \text{ mg}}\right) \left(\frac{\text{mol}}{32 \text{ gm}}\right) \left(\frac{1000 \text{ lit}}{\text{m}^3}\right) \left(\frac{32000 \text{ m}^3}{\text{day}}\right) = 2 \times 10^5 \frac{\text{mol } O_2}{\text{day}}$

Thus the rate of reaction

$$\left. \begin{aligned} \frac{2.0 \times 10^5 \text{ mol } O_2 / \text{day}}{10667 \text{ m}^3} &= 18.75 \text{ mol/m}^3 \cdot \text{day} \\ &= 2.17 \times 10^{-4} \text{ mol/m}^3 \cdot \text{s} \end{aligned} \right\} \leftarrow$$

- 1.3 Find $-r'_{C_{20}H_{42}}$ and $-r^M_{C_{20}H_{42}}$... evaluate terms

$$V_{\text{cat}} = \frac{50000 \text{ kg}}{800 \text{ kg/m}^3} = 62.5 \text{ m}^3 \text{ of catalyst}$$

$$W_{\text{cat}} = 50000 \text{ kg}$$

$$\text{MW}_{C_{20}H_{42}} = [20(12) + 42(1)] \frac{1}{1000} = 0.282 \text{ kg/mol}$$

$$F_{\text{feed}} = (6000 \text{ m}^3/\text{day})(900 \text{ kg/m}^3) = 5400000 \text{ kg/day}$$

So $-\frac{dN_{C_{20}H_{42}}}{dt} = \left(\frac{5400000 \text{ kg/day}}{0.282 \text{ kg/day}}\right) \left(\frac{\text{day}}{24(3600) \text{ s}}\right) (0.6) = 133 \text{ mol reacted/s}$

Thus the rate of disappearance of $C_{20}H_{42}$

$$\left. \begin{aligned} -r' &= \frac{1}{W_{\text{cat}}} \frac{dN}{dt} = \frac{1}{50000} (133) = 0.0027 \text{ mol/kg cat} \cdot \text{s} \\ -r^M &= \frac{1}{V_{\text{cat}}} \frac{dN}{dt} = \frac{1}{62.5} (133) = 2.13 \text{ mol/m}^3 \text{ cat} \cdot \text{s} \end{aligned} \right\} \leftarrow$$

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Koretsky,2012-12-17 Koretsky helps students understand and visualize thermodynamics through a qualitative discussion of the role of molecular interactions and a highly visual presentation of the material By showing how principles of thermodynamics relate to molecular concepts learned in prior courses Engineering and Chemical Thermodynamics 2e helps students construct new knowledge on a solid conceptual foundation Engineering and Chemical Thermodynamics 2e is designed for Thermodynamics I and Thermodynamics II courses taught out of the Chemical Engineering department to Chemical Engineering majors Specifically designed to accommodate students with different learning styles this text helps establish a solid foundation in engineering and chemical thermodynamics Clear conceptual development worked out examples and numerous end of chapter problems promote deep learning of thermodynamics and teach students how to apply thermodynamics to real world engineering problems

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