

Question	Answer	Marks	Guidance
5 (a)	<p><math>F_2</math> forces mark  <math>F_2</math> has van der Waals' (forces)  <b>OR</b>  <math>F_2</math> has induced dipole attractions <b>OR</b> interactions  <b>OR</b>  <math>F_2</math> has temporary <b>OR</b> instantaneous dipole(-dipole) attraction <b>OR</b> interactions ✓</p> <p>HCl forces mark  HCl has <b>permanent</b> dipole(-dipole) attractions <b>OR</b> interactions ✓</p> <p>Comparison of strength of forces between molecules mark  intermolecular force in HCl is stronger than that in <math>F_2</math>  <b>OR</b>  permanent dipoles are stronger (than induced dipoles) ✓</p> <p>Boiling point mark  more energy is required to break stronger (intermolecular) forces ✓</p>	4	<p><b>Use annotations ie ticks crosses ECF ^ etc for this part</b></p> <p><b>ALLOW</b> vdWs for van der Waals'  <b>IGNORE</b> <math>F_2</math> has covalent bond for this mark  <b>IGNORE</b> <math>F_2</math> has 'intermolecular forces'</p> <p><i>Quality of written communication: 'dipole(s)' spelled correctly and used in context for the second marking point</i>  <b>IGNORE</b> HCl has 'intermolecular forces'  <b>IGNORE</b> van der Waals' forces in HCl/  <b>DO NOT ALLOW</b> hydrogen bonding  <b>DO NOT ALLOW</b> ionic bonding</p> <p>Look for strength of force comparison anywhere in the answer  <b>ALLOW</b> ECF for hydrogen bonding in HCl being stronger than the stated intermolecular forces in <math>F_2</math>  <b>BUT DO NOT ALLOW</b> this mark if HCl or <math>F_2</math> has covalent bonds broken <b>OR</b> if HCl has ionic bonds broken (the question asks for forces between molecules)  <b>IGNORE</b> HCl has stronger van der Waals' (forces) than <math>F_2</math> (as they both have the same number of electrons)</p> <p><b>DO NOT ALLOW</b> fourth mark if covalent bonds are broken in HCl or <math>F_2</math> <b>OR</b> if ionic bonds are broken in HCl</p> <p><b>IGNORE</b> 'heat' but <b>ALLOW</b> 'heat energy'</p>

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