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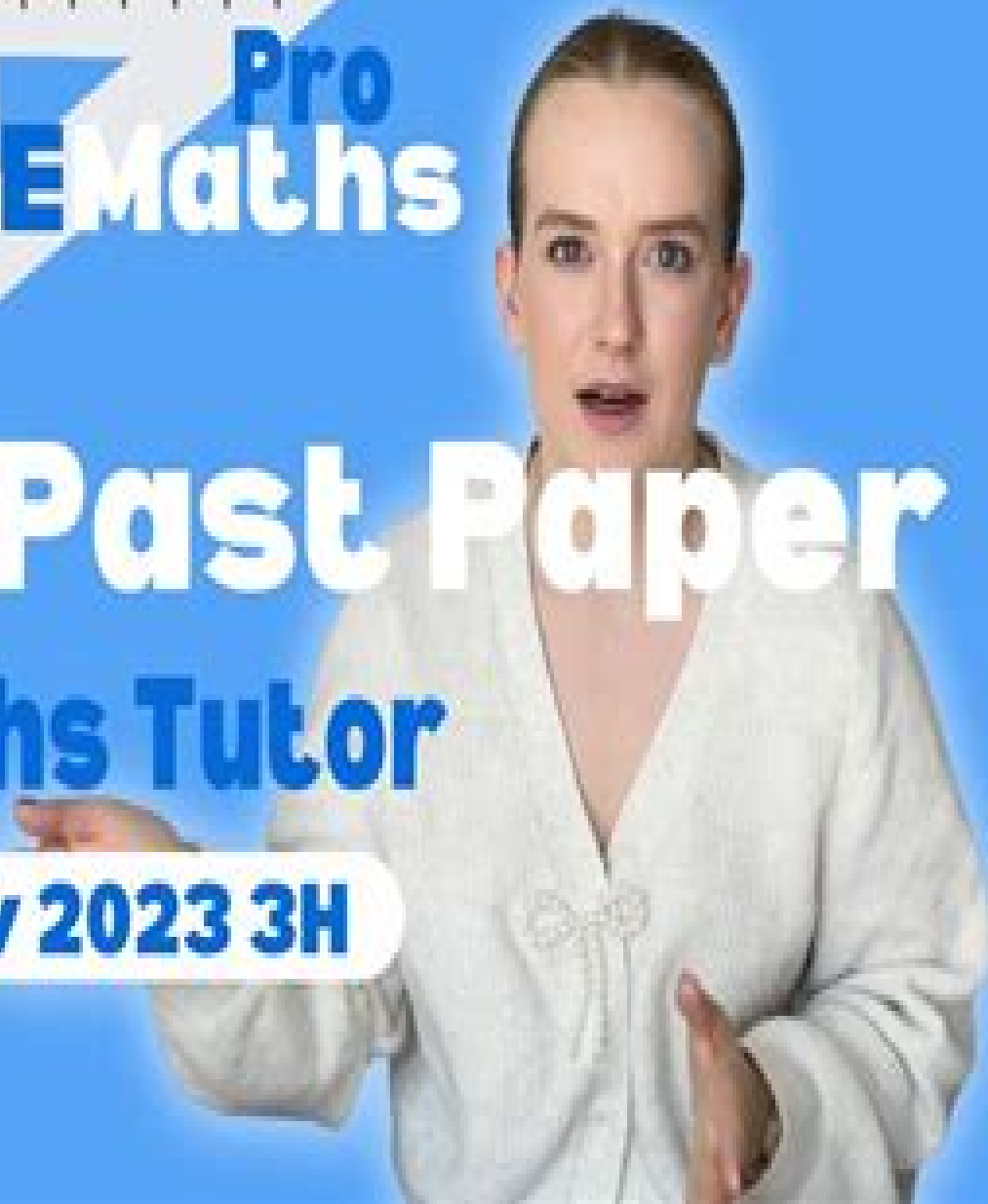
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**Léo Pomar**

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yield 3co 7h 2 c 3 h 8 3h 2 o

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web a write a balanced equation for the reaction al + o<sub>2</sub> → al<sub>2</sub>O<sub>3</sub> 4 al + 3 o<sub>2</sub> → 2 al<sub>2</sub>O<sub>3</sub> b determine the theoretical yield of al<sub>2</sub>O<sub>3</sub> 160 0 g al x 1 mol al / 27 g al x 2 mol al<sub>2</sub>O<sub>3</sub> / 3 mol al = 101 964 g al<sub>2</sub>O<sub>3</sub> 302 3 g al<sub>2</sub>O<sub>3</sub> 26 982 g al 4 mol al / 1 mol al<sub>2</sub>O<sub>3</sub> c determine the percent yield yield actual x 100 260 0 g / 302 3 g x 100 = 86 01 theoretical 302 3 g

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yield theoretical yield actual yield 100 based on this definition we would expect a percent yield

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