

**Unit 1 Progress Check: MCQ**

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1. Which of the following numerical expressions gives the number of moles in 5.0 g of  $\text{CaO}$ ?

(A)  $5.0 \text{ g} \times 56 \text{ g/mol}$

(B)  $\frac{5.0 \text{ g}}{56 \text{ g/mol}}$

(C)  $\frac{56 \text{ g/mol}}{5.0 \text{ g}}$

(D)  $\frac{1}{5.0 \text{ g}} \times \frac{1}{56 \text{ g/mol}}$



2. In a lab a student is given a 21 g sample of pure  $\text{Cu}$  metal. Which of the following pieces of information is most useful for determining the number of  $\text{Cu}$  atoms in the sample? Assume that the pressure and temperature in the lab are 1.0 atm and  $25^\circ\text{C}$ .

(A) The molar mass of  $\text{Cu}$

(B) The density of  $\text{Cu}$  at  $25^\circ\text{C}$

(C) The volume of the  $\text{Cu}$  sample

(D) The ratio of the two main isotopes found in pure  $\text{Cu}$



3. A 1.0 mol sample of which of the following compounds has the greatest mass?



# Novanet Chemistry Answers

**SJ Ball**



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