

Chapter 7.1

Solutions

- Homogeneous mixtures of two or more substances in which the components are present as atoms, molecules, or ions
- Particles in liquid solutions are:
 - Too small to reflect light, thus solutions are transparent
 - In constant motion
 - Not settled by the influence of gravity
- Some solutions are colored
- Solvent of a solution
 - The substance present in the largest amount of the solution
- Solute of a solution
 - Any substance present in an amount less than that of the solvent
 - A solution may contain more than one solute
- Characteristics of a liquid solution
 - It has uniform composition
 - It cannot be separated by filtration
 - It can be separated into pure components
- Physical states of solutions
 - The physical state of a solution of a solid, liquid, or gas, is usually the same as the physical state of the solvent

TABLE 7.1 Solutions in various states

Solution	Solution state	Solvent	Solute
Salt water	Liquid	Water (liquid)	Sodium chloride (solid)
Alcoholic beverage	Liquid	Water (liquid)	Alcohol (liquid)
Carbonated water	Liquid	Water (liquid)	Carbon dioxide (gas)
Gold alloy (jewelry)	Solid	Gold (solid)	Copper (solid)
Gold amalgam	Solid	Gold (solid)	Mercury (liquid)
Hydrogen in palladium	Solid	Palladium (solid)	Hydrogen (gas)
Air	Gaseous	Nitrogen (gas)	Oxygen (gas)
Mixed oxygen	Gaseous	Oxygen (gas)	Water (liquid)
Camphor in nitrogen	Gaseous	Nitrogen (gas)	Camphor (solid)

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- Solubility
 - The solubility of a solute is the maximum amount of the solute that can be dissolved in a specific amount of solvent under specific conditions of temperature and pressure
 - Solute solubility

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Table of Contents Noun Chm 103

1. Understanding the eBook Noun Chm 103
 - The Rise of Digital Reading Noun Chm 103
 - Advantages of eBooks Over Traditional Books
2. Identifying Noun Chm 103
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Noun Chm 103
 - User-Friendly Interface
4. Exploring eBook Recommendations from Noun Chm 103
 - Personalized Recommendations
 - Noun Chm 103 User Reviews and Ratings
 - Noun Chm 103 and Bestseller Lists
5. Accessing Noun Chm 103 Free and Paid eBooks
 - Noun Chm 103 Public Domain eBooks
 - Noun Chm 103 eBook Subscription Services
 - Noun Chm 103 Budget-Friendly Options
6. Navigating Noun Chm 103 eBook Formats

- ePub, PDF, MOBI, and More
- Noun Chm 103 Compatibility with Devices
- Noun Chm 103 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Noun Chm 103
 - Highlighting and Note-Taking Noun Chm 103
 - Interactive Elements Noun Chm 103
- 8. Staying Engaged with Noun Chm 103
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Noun Chm 103
- 9. Balancing eBooks and Physical Books Noun Chm 103
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Noun Chm 103
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Noun Chm 103
 - Setting Reading Goals Noun Chm 103
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Noun Chm 103
 - Fact-Checking eBook Content of Noun Chm 103
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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